

Gas Laws HOMEWORK

Lesson 1 Homework

1. Define pressure. *Force that a gas exerts on its container*

2. Convert the following:

a. 750 mmHg to atm

b. 550 mmHg to atm

c. 2.5 atm to torr

d. 2.1×10^3 Pa to atm

e. 300 Pa to atm

$$750 \text{ mmHg} \times (1 \text{ atm} / 760 \text{ mmHg}) = 0.9868 \text{ atm}$$

$$550 \text{ mmHg} \times (1 \text{ atm} / 760 \text{ mmHg}) = 0.72368 \text{ atm}$$

$$2.5 \text{ atm} \times (760 \text{ torr} / 1 \text{ atm}) = 1900 \text{ torr}$$

$$2.1 \times 10^3 \text{ Pa} \times (1 \text{ atm} / 101325 \text{ Pa}) = 0.02072 \text{ atm}$$

$$300 \text{ Pa} \times (1 \text{ atm} / 101325 \text{ Pa}) = 2.96 \times 10^{-3} \text{ atm}$$

3. Calculate the atmospheric pressure in kPa on a day when the pressure is 750 mm.

$$750 \text{ mm} \times (101.325 \text{ kPa} / 760 \text{ mm}) = 100.867 \text{ kPa} \rightarrow 100.9 \text{ kPa}$$

4. Calculate the pressure of a sample of gas (in mmHg) inside an open-end manometer if the mercury column is 35 mm higher on the side of the closed tube than on the atmosphere. The atmospheric pressure is 750 mm. Draw a picture to illustrate your answer.

$$750 \text{ mm} + (170 \text{ mmHg} / 1 \text{ mm}) = 767 \text{ mmHg}$$

750 mm mercury is 750 mm Hg because the atmospheric pressure, then it's 170 mmHg because that's the height.

$$767 \text{ mmHg} \times (1 \text{ mm} / 1 \text{ mm}) = 767 \text{ mmHg}$$

Since pressure should decrease a centimeter with the height of the mercury, then there the atmospheric level of 750 mm.

5. If a gas at a pressure of 150 mm in a 1.0 L container has its pressure increased to 400 mm, what is the new volume of the container? (Assume temperature is constant).

$$P_1 V_1 = P_2 V_2$$

$$(150 \text{ mm} \times 1.0 \text{ L}) = (400 \text{ mm} \times V_2)$$

$$150 = 400 V_2 \rightarrow 0.375 \text{ L}$$

6. What is the new pressure if a 5.00 closed container with a pressure of 2.5 atm is increased to 8.0 L?

$$P_1 V_1 = P_2 V_2$$

$$(2.5 \text{ atm} \times 5.00 \text{ L}) = (P_2 \times 8.0 \text{ L})$$

$$12.5 = 8.0 P_2 \rightarrow 1.56 \text{ atm}$$

7. When 0.5 L of oxygen and 4.0 L of Nitrogen both at 0°C and 1.0 atm, are pumped into an evacuated 5.0 L rigid container, what is the final pressure in the container at 0°C? *Since temperature doesn't change, you can leave it out of the equation.*

So final pressure of Nitrogen:

$$P_1 V_1 = P_2 V_2$$

$$(1.0 \text{ atm} \times 4.0 \text{ L}) = (P_2 \times 5.0 \text{ L})$$

$$4.0 = 5.0 P_2$$

So final pressure of Nitrogen:

$$P_1 V_1 = P_2 V_2$$

$$(1.0 \text{ atm} \times 0.5 \text{ L}) = (P_2 \times 5.0 \text{ L})$$

$$0.5 = 5.0 P_2$$

$$\text{Total Pressure} = P_1 + P_2 + P_3$$

$$1.0 = 1.6 + 1.6 + 1.6 \rightarrow 3.2 \text{ atm}$$

Atmosphere would combine the two values

$$P_1 V_1 = P_2 V_2$$

$$(1.0 \text{ atm} \times 4.0 \text{ L}) = (P_2 \times 5.0 \text{ L})$$

$$4.0 = 5.0 P_2 \rightarrow 0.8 \text{ atm}$$

Then make another volume and make are directly proportional

8. A sealed flask contains a gas with pressure P. If the number of moles of the gas is doubled, the absolute temperature is doubled and the volume of the container is quadrupled, what is the new pressure in terms of P?

Doubling the moles doubles the pressure

$$P_1 V_1 = P_2 V_2$$

Quadrupling the volume cuts the pressure in 1/4

$$P_1 V_1 = P_2 V_2$$

moles do no overall change in total P

9. If 4 moles of gas are added to a container that already holds 1 mole of gas, how will the pressure change within the container?

$$P_1 V_1 = P_2 V_2$$

Going from 1 mole to 5 moles results in an increase of pressure by 4 times

$$P_1 V_1 = P_2 V_2$$

10. If a balloon containing 1000 L of gas at 40.°C and 760 mmHg rises to an altitude where the pressure is 500 mmHg and the temperature is 10.0°C, what is the new volume of the balloon (in L)?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \times \frac{1 \text{ L}}{1 \text{ L}}$$

$$(760 \text{ mmHg} \times 1000 \text{ L}) / (313 \text{ K}) = (500 \text{ mmHg} \times V_2) / (283 \text{ K})$$

$$2400000 = 1.76 V_2$$

$$V_2 = 13636.36 \rightarrow 13636 \text{ L}$$

11. A sample of nitrogen gas has a pressure of 87.5 mmHg in a 500. mL flask. What is the pressure of this gas sample when it is transferred to a 225.0 mL flask at the same temperature?

$$P_1 V_1 = P_2 V_2$$

$$(87.5 \text{ mmHg} \times 500 \text{ mL}) = (P_2 \times 225 \text{ mL})$$

$$43750 = 225 P_2$$

$$P_2 = 194.44 \rightarrow 194 \text{ mmHg}$$

12. You have 4.5 L of NO_2 at a temperature of 23.0°C. What volume would the NO_2 occupy at 47°C? (Assume the pressure is constant).

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \times \frac{1 \text{ L}}{1 \text{ L}}$$

$$(4.5 \text{ L} \times 295 \text{ K}) / (273 \text{ K}) = (V_2 \times 320 \text{ K}) / (273 \text{ K})$$

$$1327.5 = 320 V_2$$

$$V_2 = 4.15 \rightarrow 4.15 \text{ L}$$

13. You have a sample of a gas in a flask with a volume of 240 mL. At 35.0°C the pressure of the gas is 540 mmHg. If you decrease the temperature to -15.0°C, what is the gas pressure at the lower temperature? *Since volume doesn't change, you can leave it out of the equation.*

$$P_1 / T_1 = P_2 / T_2$$

$$(540 \text{ mmHg} \times 310 \text{ K}) / (273 \text{ K}) = (P_2 \times 240 \text{ mL}) / (273 \text{ K})$$

$$540 \times 310 = 240 P_2$$

$$P_2 = 708.75 \rightarrow 709 \text{ mmHg}$$

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