

(Solution Stoichiometry, Precipitation Reactions, Acid/Base Reactions, Multi-Concept Reaction Stoichiometry)

Solution Stoichiometry: Molarity, Dilutions and Working with Density

1. Calculate the molarity of each of the following solutions:

a. 2.026 g of FeCl_3 in 0.1250 L of solution

$$\frac{2.026 \text{ g FeCl}_3}{0.1250 \text{ L}} \times \frac{1 \text{ mole FeCl}_3}{162.2 \text{ g}} = 0.09993 \text{ M}$$

b. 0.001 mg of Cd^{2+} in 0.100 L, the maximum permissible concentration of cadmium in drinking water

$$\frac{0.001 \text{ mg Cd}^{2+}}{0.100 \text{ L}} \times \frac{1 \text{ g}}{1000 \text{ mg}} \times \frac{1 \text{ mole Cd}^{2+}}{112.41 \text{ g Cd}^{2+}} = \sim 9 \times 10^{-8} \text{ M}$$

2. Convert 1b to ppm and ppb.

$$\text{ppm} = \text{Parts per million} \quad \text{i.e.} \quad \frac{\text{g}}{10^6 \text{ g}} \text{ or } \frac{\text{g}}{10^6 \text{ mL}} \text{ or } \frac{\text{mg}}{\text{L}}$$

$$\text{ppb} = \text{Parts per billion} \quad \text{i.e.} \quad \frac{\text{g}}{10^9 \text{ g}} \text{ or } \frac{\text{g}}{10^9 \text{ mL}} \text{ or } \frac{\mu\text{g}}{\text{L}}$$

$$\frac{0.001 \text{ mg Cd}^{2+}}{0.100 \text{ L}} \times \frac{1 \text{ ppm}}{\text{mg/L}} = 0.01 \text{ ppm}$$

$$\frac{0.001 \text{ mg Cd}^{2+}}{0.100 \text{ L}} \times \frac{1 \text{ g}}{1000 \text{ mg}} \times \frac{10^6 \mu\text{g}}{1 \text{ g}} \times \frac{1 \text{ ppb}}{\mu\text{g/L}} = 10 \text{ ppb}$$

Gas Law Stoichiometry Worksheet Answers

JS Bruner

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