WS Gas Laws - KEY

Boyle's Law $P_1 \times V_1 = P_2 \times V_2$ At constant temperature Combined Gas Law Charles's Law $V_1 - V_2$ At constant pressure T_{ij} The Temperature in Kelvip C + 273 = K $\mathbf{T}_{\mathbf{i}}$ Lussac's Law At constant volume $P_{\perp} =$ T_{A} STP: 273 K & 1.00 atm Temperature in Kelvin

Concepts: (circle one choice)

- If the volume of a container of gas is reduced, what will happen to the pressure inside the container? increase or decrease
- What happens to the volume of a gas when it is cooled to a lower temperature? increase or decrease
- What happens to the pressure in a rigid container when it is slowly warmed? increase or decrease

Sample Calculations: (Remember: <u>ALL</u> temperatures <u>MUST</u> be in <u>K</u>, <u>NOT</u> °C)

 A sample of gas at 240 K and 670 mmHg occupies a 0.128 L volume. What volume will the gas occupy at 198 K if the pressure remains constant?

$$\frac{\mathbf{X}_1 \, \mathbf{V}_1 = \mathbf{X}_2 \, \mathbf{V}_2}{\mathbf{T}_1 \quad \mathbf{T}_2} \qquad \frac{(0.128 \, \mathbf{L}) = \mathbf{V}_2}{(240 \, \mathrm{K})} \qquad \frac{(0.128)(198)}{(198 \, \mathrm{K})} = \mathbf{V}_2 = 0.106 \, \mathbf{L}$$

 A sample of gas is in a steel container at -75.0°C and 1.48 atm. At what temperature will the sample have a pressure of 7.35 atm?

3. In an airplane, a gas sample occurs at a volume of 1.50 L at 760 mmHg. Suppose, while flying, the airplane loses pressure and the volume of the gas increases to 11.40 L. What is the pressure in the airplane if the temperature is constant?

$$P_1 V_1 = P_2 V_2$$
 (760 mmHg)(1.50 L) = P_2 (11.40 L) (760)(1.50) = P_2 = 100 mmHg
 $X_1 X_2$ (11.40)

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