

1. The theory that explains the behaviour of gases at the molecular level is called the \_\_\_\_\_ theory. \_\_\_\_\_ reflects its kinetic assumptions about a theoretical gas which is confined in a container \_\_\_\_\_.
2. Gases behave most free-like gas behaviour under conditions of very low \_\_\_\_\_ and very high \_\_\_\_\_.  
 The molecules of an ideal gas display the \_\_\_\_\_ as \_\_\_\_\_ for equilibrium.  
 Under ordinary conditions, no ideal gas behaves exactly as \_\_\_\_\_  
 which explains why gases are so easily compressed.  
 Ideal gas particles move in \_\_\_\_\_ lines until they collide with each other or with the walls of their container.  
 The collisions between the molecules of an ideal gas are completely \_\_\_\_\_.  
 The average kinetic energy of the molecules of an ideal gas is \_\_\_\_\_ proportional to the \_\_\_\_\_ temperature of the gas.
3. A gas thermometer on the walls of its container behaves gas molecules \_\_\_\_\_ with the walls of the container. So, the pressure exerted by a gas depends on how fast they:
  - (a) \_\_\_\_\_
  - (b) \_\_\_\_\_
4. To measure gas pressure an instrument called a \_\_\_\_\_ is used.
5. The earth's atmosphere has weight, which creates \_\_\_\_\_.
6. The instrument used to measure atmospheric pressure is the \_\_\_\_\_.

# Gas Law Note Taking Guide 901 Key

**Hongru Du**



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