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Determination of the Empirical Formula of Silver Oxide

AP Chemistry Laboratory #1

Catalog No. AP8813

Publication No. 10528A

Introduction

There is an official database that keeps track of the known chemical compounds that exist in nature or have been synthesized in the lab. The database, called the chemical abstracts database, is updated daily. Currently, over 20 million different inorganic and organic compounds have been recognized. Twenty million compounds—how is it possible to identify so many different compounds and tell them all apart?

Concepts

- Percent composition
- Empirical formula
- Law of conservation of mass
- Molecular formula
- Percent yield

Background

The composition of a chemical compound—what it is made of—can be described at least three different ways. The *percent composition* gives the percent by mass of each element in the compound and is the simplest way experimentally to describe the composition of a substance. According to the law of definite proportions, which was first formulated in the early 1800s by Joseph Proust, the elements in a given compound are always present in the same proportion by mass, regardless of the source of the compound or how it is prepared. Calcium carbonate, for example, contains calcium, carbon, and oxygen. It is present in eggshells and seashells, chalk and limestone, minerals and pearls. Whether the calcium carbonate comes from a mineral supplement on a drugstore shelf or from seashells on the ocean shore, the mass percentage of the three elements is always the same: 40% calcium, 12% carbon, and 48% oxygen.

The percent composition of a compound tells us what elements are present in the compound and their mass ratio. In terms of understanding how elements come together to make a new compound, however, it is more interesting and more informative to know how many atoms of each kind of element combine with one another. Since all the atoms of a given element in a compound have the

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Gourmet Lab Sarah Reeves Young, 2011 Hands on inquiry based and relevant to every student's life Gourmet Lab serves up a full menu of activities for science teachers of grades 6-12 This collection of 15 hands on experiments each of which includes a full set of both student and teacher pages challenges students to take on the role of scientist and chef as they boil bake and toast their way to better understanding of science concepts from chemistry biology and physics By cooking edible items such as pancakes and butterscotch students have the opportunity to learn about physical changes in states of matter acids and bases biochemistry and molecular structure The Teacher pages include Standards addressed in each lab a vocabulary list safety protocols materials required procedures data analysis student questions answer key and conclusions and connections to spur wrap up class discussions Cross curricular notes are also included to highlight the lesson's connection to subjects such as math and literacy Finally optional extensions for both middle school and high school levels detail how to explore each concept further What better topic than food to engage students to explore science in the natural world

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Regions and Industries Pat Hudson, 1989-10-26 In this book a team of distinguished historians contend that

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