

Lab 7: Flame Test Lab

Prelab Reading

The normal electron configuration of atoms or ions of an element is known as the "ground state". In this most stable energy state all of the electrons are in the lowest energy levels available. When atoms or ions in the ground state are heated to high temperatures some electrons may absorb energy and "jump" to a higher energy level. The element is then said to be in an "excited state". This excited configuration is unstable and the electrons will "fall back" to their normal positions of lower energy. As the electrons return to their normal levels the energy that was absorbed is now emitted in the form of electromagnetic radiation. Some of this radiation is in the form of visible light.

The color of this light can be used as a means of identifying the elements involved. This analysis is known as a flame test. Only metals, with their loosely held electrons, are excited by the flame of a bunsen burner. Thus flame tests are useful in the identification of metallic ions. Many metallic ions exhibit characteristic colors when vaporized by the burner flame. In this experiment characteristic colors of several different metallic ions will be observed and unknown ions will be identified by means of the flame test.

Before beginning the pre-lab questions below, be sure to read the procedure on the next page.

Pre-Lab Questions:

1. What is the difference between ground state and an excited state?

Ground state: all electrons are at the lowest possible energy levels.

Excited state: Atoms/ions absorb energy and go to a higher energy.

2. What does the word "emit" mean?

Emit is defined as being produced (i.e. the energy that was absorbed is produced in the form of electromagnetic radiation).

3. In this experiment, where are the atoms getting their excess energy from?

The electrons are getting the excess energy from a bunsen burner.

4. Why do different atoms emit different colors of light?

Atoms emit different colors because of their chemical makeup.

5. Why is it necessary to clean the nichrome wires between each flame test?

It is necessary to clean the wires between the flame tests because if you don't, the atoms might not display their characteristic color because they were mixed with another atom.

Flame Test Lab Answers

R Barnett



Flame Test Lab Answers:

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