

3

Classification of Elements and Periodicity in Properties

Multiple Choice Questions (MCQs)

Q. 1 Consider the isoelectronic species, Na^+ , Mg^{2+} , F^- and O^{2-} . The correct order of increasing length of their radii is

- (a) $\text{F}^- < \text{O}^{2-} < \text{Mg}^{2+} < \text{Na}^+$ (b) $\text{Mg}^{2+} < \text{Na}^+ < \text{F}^- < \text{O}^{2-}$
(c) $\text{O}^{2-} < \text{F}^- < \text{Na}^+ < \text{Mg}^{2+}$ (d) $\text{O}^{2-} < \text{F}^- < \text{Mg}^{2+} < \text{Na}^+$

Ans. (b) In case of isoelectronic species ionic radii $\propto \frac{1}{\text{atomic number}}$
The ionic radii increases as the positive charge decreases or the negative charge increases.

| | | | | |
|---------------|------------------|-----------------|----------------|-------------------|
| Ion | Mg^{2+} | $< \text{Na}^+$ | $< \text{F}^-$ | $< \text{O}^{2-}$ |
| Atomic number | (12) | (11) | (9) | (8) |

Q. 2 Which of the following is not an actinoid?

- (a) Curium ($Z = 96$) (b) Californium ($Z = 98$)
(c) Uranium ($Z = 92$) (d) Terbium ($Z = 65$)

Ans. (d) Elements with atomic number, $Z = 90$ to 103 are called actinoids. Thus, terbium ($Z = 65$) is not an actinoid. Terbium belongs to lanthanoids.

Q. 3 The order of screening effect of electrons of s , p , d and f orbitals of a given shell of an atom on its outer shell electrons is

- (a) $s > p > d > f$ (b) $f > d > p > s$ (c) $p < d < s > f$ (d) $f > p > s > d$

Thinking Process

To solve this question, keep in mind that shielding effect represents the repulsive force felt by the valence shell from the electrons present in the inner shells.

Ans. (a) For the same shell, penetration of electrons in the order $s > p > d > f$.

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