

BASIC CHEMISTRY

CHEMISTRY

- ✓ Medieval Latin "Alchimia"/ Arabic "Alkimya", the chemistry
- ✓ The study of the composition and structure of the changes in matter
- ✓ It is often called the central science because knowledge in this field is a prerequisite for all the other physical and biological sciences.

HISTORICAL DEVELOPMENT OF THE ATOMIC STRUCTURE

ATOM

- ✓ It is the smallest possible unit into which matter can be divided, while still maintaining its properties
 - ✓ Building blocks of matter
 - ✓ It has protons (+), neutrons (0), and electrons (-)
1. **Democritus** was a greek philosopher who first introduced the concept of atoms. He believed that matter is comprises of tiny indivisible particles known as atomos.
 2. **Dalton's Atomic Theory**
 - a. All matter is made up of indivisible particles known as atoms.
 - b. All atoms of the same elements are alike.
 - c. Compounds are formed when atoms of different elements combine in certain proportions.
 - d. atoms are only rearranged but neither created nor destroyed in chemical reactions.
 3. **Law of Definite Proportion** – a given compound always contains the same elements in exactly the same proportion by mass.

H_2O - ratio of two hydrogen atom with one oxygen atom only. If it becomes ratio of two hydrogen atom with two oxygen atom, it is not already water.
 4. **Law of Multiple Proportions** – when two elements can combine to form more than one type of compound, the masses of one element that combine with a fixed mass of the other elements are in a ratio of small whole numbers.

$\text{CaO} + \text{CO}_2 \rightarrow \text{CaCO}_3$
 5. **Law of Conservation of Mass** states that matter can neither be created nor destroyed. Matter is just transformed.
 6. The discovery of the different sub-atomic particles:
 - i. **J.J. Thompson** determined the charge to mass ratio of the electron using a cathode ray tube.

Example Basic Chemistry Study Guide

Jill Meryl Levy



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