

# Atoms and Bonds Study Guide

## Section 1 Fill in the Table

| Compound/Molecule             | Metal + Nonmetal or<br>Nonmetal + Nonmetal? | Ionic or Covalent Bond? | e <sup>-</sup> shared or<br>transferred? |
|-------------------------------|---|-------------------------|--|
| NaCl                          |   |                         |  |
| C <sub>2</sub> H <sub>6</sub> |   |                         |  |
| PF <sub>5</sub>               |   |                         |  |
| MgO                           |   |                         |  |

## Section 2 Periodic Symbol Analysis

|        |
|--------|
| 74     |
| W      |
| 183.84 |

Element name \_\_\_\_\_  
 Name English or Latin in origin? \_\_\_\_\_  
 Atomic number \_\_\_\_\_ Number of protons \_\_\_\_\_  
 Atomic mass \_\_\_\_\_ Number of electrons \_\_\_\_\_  
 Number of neutrons (Mass - # protons) \_\_\_\_\_  
 Metal, nonmetal, or noble gas? \_\_\_\_\_

**Section 3 Properties of Bonds** Write "I" if the property describes an ionic bond, "C" if the property describes a covalent bond, and "B" if the property describes both kinds of bonds.

\_\_\_\_\_ High melting points      \_\_\_\_\_ Weaker bond      \_\_\_\_\_ Electrons shared  
 \_\_\_\_\_ Conducts electricity in water      \_\_\_\_\_ Helps most atoms achieve the octet of a noble gas  
 \_\_\_\_\_ Polar or nonpolar molecules      \_\_\_\_\_ Attraction of opposite charges  
 \_\_\_\_\_ Poor conductors of electricity      \_\_\_\_\_ Crystal shape      \_\_\_\_\_ Low melting points  
 \_\_\_\_\_ Attraction between a cation and an anion      \_\_\_\_\_ Formed or broken in a chemical reaction

**Section 4 Formula Writing** Write the formula for each of the compounds below. Be sure to balance the charges with subscripts. [Hint: Parenthesized Roman numerals are the charges of transition metals.]

- Copper (II) sulfate \_\_\_\_\_
- Mercury (I) oxide \_\_\_\_\_
- Mercury (II) nitrate \_\_\_\_\_
- Beryllium nitrate \_\_\_\_\_
- Aluminum oxide \_\_\_\_\_
- Magnesium hydroxide \_\_\_\_\_

Sulfate = SO<sub>4</sub><sup>2-</sup>Oxide = O<sup>2-</sup>Nitrate = NO<sub>3</sub><sup>-</sup>Hydroxide = OH<sup>-</sup>

# Evidence For Atoms Webquest Teacher Guide

**Tina Youngman**



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## **Table of Contents Evidence For Atoms Webquest Teacher Guide**

1. Understanding the eBook Evidence For Atoms Webquest Teacher Guide
  - The Rise of Digital Reading Evidence For Atoms Webquest Teacher Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Evidence For Atoms Webquest Teacher Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Evidence For Atoms Webquest Teacher Guide
  - User-Friendly Interface

4. Exploring eBook Recommendations from Evidence For Atoms Webquest Teacher Guide
  - Personalized Recommendations
  - Evidence For Atoms Webquest Teacher Guide User Reviews and Ratings
  - Evidence For Atoms Webquest Teacher Guide and Bestseller Lists
5. Accessing Evidence For Atoms Webquest Teacher Guide Free and Paid eBooks
  - Evidence For Atoms Webquest Teacher Guide Public Domain eBooks
  - Evidence For Atoms Webquest Teacher Guide eBook Subscription Services
  - Evidence For Atoms Webquest Teacher Guide Budget-Friendly Options
6. Navigating Evidence For Atoms Webquest Teacher Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Evidence For Atoms Webquest Teacher Guide Compatibility with Devices
  - Evidence For Atoms Webquest Teacher Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Evidence For Atoms Webquest Teacher Guide
  - Highlighting and Note-Taking Evidence For Atoms Webquest Teacher Guide
  - Interactive Elements Evidence For Atoms Webquest Teacher Guide
8. Staying Engaged with Evidence For Atoms Webquest Teacher Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Evidence For Atoms Webquest Teacher Guide
9. Balancing eBooks and Physical Books Evidence For Atoms Webquest Teacher Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Evidence For Atoms Webquest Teacher Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Evidence For Atoms Webquest Teacher Guide
  - Setting Reading Goals Evidence For Atoms Webquest Teacher Guide
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Evidence For Atoms Webquest Teacher Guide
  - Fact-Checking eBook Content of Evidence For Atoms Webquest Teacher Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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