

15. There are 618 books in the classroom library. Complete the chart to show 618 rounded to the nearest 10.

Hundreds	Tens	Ones

16. Select all expressions that are equal to 384.

- A.  $203 + 181$
- B.  $168 + 206$
- C.  $287 + 97$
- D.  $609 - 225$
- E.  $523 - 149$

17. There are 4 tables in a classroom. Each table has 5 students. Each student has 6 markers. How many markers are in the classroom?

18. A store has 20 crates of apples. Each crate holds 6 bags. Each bag holds 8 apples. What is the total number of apples in the store?

19. Select all expressions that have a product of 240.

- A.  $4 \times 60$
- B.  $8 \times 40$
- C.  $3 \times 80$
- D.  $5 \times 50$
- E.  $6 \times 40$

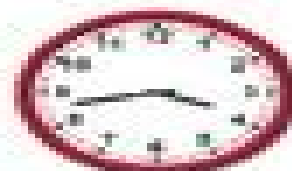
20. Abel has 518 emails in his inbox. He deletes 375 emails. How many emails remain in his inbox?

\_\_\_\_\_ emails

21. What is the sum of 235, 183, and 87?

\_\_\_\_\_

22. Kathleen checked the time on her watch when she got home from school.



Select the time that Kathleen got home from school. Select all that apply.

- A. seven-thirty minutes before four.
- B. four sixty-two
- C. three twenty-five
- D. three fifty-five
- E. eight minutes

# Everyday Math End Of Year Test

**Camilla Rothe**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{3+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{2+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- + 2\text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CN}^-$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{Au} + \text{CN}^- + 2\text{e}^-$  multiplied by a factor of 2  $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^-$  the final answer

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$\text{H}_2\text{O}$  pair to fully balance the equation in this video we'll walk through this process for the reaction between  $\text{ClO}^-$  and  $\text{Cr}(\text{OH})_4^-$  in basic solution created by jay

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web balance each half reaction in basic solution  $4\text{SO}_3^{2-} + 2\text{MnO}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{NO}_3^- \rightarrow 4\text{SO}_4^{2-} + 2\text{Mn}^{2+} + \text{Cr}^{3+} + 2\text{NO}_2$  balance each redox reaction in acid solution using the half reaction method  $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{H}_2\text{O} + 2\text{Cr}^{3+} + 10\text{H}_2\text{O}$   $11\text{PbO}_2 + \text{IO}_3^- + 2\text{H}^+ \rightarrow \text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{SO}_4$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h 2no 2 c 6 h 6 o 6 2h 2 o 2no choose 1 answer c 6 h 8 o 6 and no 2 a c 6 h 8 o 6 and no 2 no 2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and

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