

- Q6 When water was added to the anhydrate of  $\text{CuSO}_4(s)$ , did the reaction proceed in the direction of the reactants or products?

## B.2 Reaction of Copper(II) Ion and Hydroxide

	Test tube 1	Test tube 2
Initial color of $\text{CuCl}_2$ solution (1)	lighter blue liquid transparent	lighter blue liquid transparent
Initial appearance $\text{Cu}(\text{OH})_2(s)$ (2)	cloudy blue solution	cloudy blue solution
Change in appearance (3)	more cloudy solution	
Change in appearance (4)		clear lighter blue solution

- Q7 In test tube 1, did the addition of  $\text{OH}^-$  cause the reaction to go in the direction of the reactants or products? Explain.

- Q8 In test tube 2, did the addition of  $\text{H}_3\text{O}^+$  cause the reaction to go in the direction of the reactants or products? Explain.

## C. Changing Equilibrium Conditions: Le Châtelier's Principle

### C.1 Effect of Changing Reactant Concentration

Test tube	Color	Is color deeper or lighter red compared to the reference?	$[\text{FeSCN}^{2+}]$ increases or decreases?
1. Reference	blood red		
2. $\text{Fe}^{3+}$ added	dark blood red	deeper	
3. $\text{SCN}^-$ added	dark blood red	deeper	
4. $\text{Fe}^{3+}$ removed	light blood red	lighter	

# Evaluation Reaction Rates And Chemistry Equilibrium Answers

**Thomas Griffiths**



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