

Chemistry I Study Guide Chapters 1-3

Chapter 1

Key Concepts

- 1. What is a molecule?*
- 2. What are the types of molecules?*
- 3. What are the three types of bonds in molecules?*
- 4. What are the four types of molecules?*
- 5. What are the five types of molecules?*
- 6. What are the six types of molecules?*
- 7. What are the seven types of molecules?*
- 8. What are the eight types of molecules?*
- 9. What are the nine types of molecules?*
- 10. What are the 10 types of molecules?*

Important Events

- 1. In 1827, Friedrich Wöhler synthesized urea from carbons and hydrogens.*
- 2. In 1828, Justus von Liebig isolated and identified nitrogen as a component of organic compounds.*
- 3. In 1828, Michael Faraday discovered that electricity and magnetism are related.*
- 4. In 1831, Friedrich Wöhler synthesized ammonia from nitrogen and hydrogen.*
- 5. In 1835, Jean-Baptiste Dumas and Michel Eugène Chevreul identified sulfur as a component of organic compounds.*
- 6. In 1839, Friedrich Wöhler synthesized glycerin from fats and water.*
- 7. In 1841, John Dalton proposed that all elements are made of atoms of the same type.*
- 8. In 1844, Lewis Paul and John Dalton independently proposed that the elements are arranged in a periodic table.*
- 9. In 1845, Friedrich Wöhler synthesized urea from carbon and hydrogen.*
- 10. In 1851, Louis Pasteur demonstrated that fermentation is caused by bacteria.*

Notable Figures

- 1. Justus von Liebig*
- 2. Friedrich Wöhler*
- 3. Michael Faraday*
- 4. Jean-Baptiste Dumas*
- 5. John Dalton*
- 6. Lewis Paul*
- 7. Friedrich Wöhler*

Intro To Chemistry Study Guide

Steven S. Zumdahl, Donald J. DeCoste



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Introduction to Chemistry, Study Guide T. R. Dickson, 1991-01-16 A thorough grounding in the fundamentals in any discipline is crucial Chemistry is no exception This Sixth Edition reflects the author's continuing emphasis on the introductory chemistry course as the most important of all that chemistry students will take Continual revision is necessary and it has been rewritten to improve clarity and update the material presented in previous editions Based on feedback from students and professors alike the material is now offered in a different sequence and end of chapter questions have been revised and expanded It continues to provide a dynamic interesting and relevant view of the subject matter and the environment with the major theme being the introduction explanation and illustration of the problem solving methods of beginning chemistry Approaches to solving chemical problems by reasoning and the unit equation factor label or dimensional analysis methods are explained in detail with numerous examples incorporated Relevant examples analogies and special topics continue to reinforce introduce and illustrate chemical concepts

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