

# Intro to Acids and Bases

Name: \_\_\_\_\_  
 Class: \_\_\_\_\_  
 Date: \_\_\_\_\_

## Define each term: Definitions of acids and bases:

### Arrhenius-definitions:

- acid substances that release dissolved cations hydrogen ( $H^+$ ) (acids  $H^+$  cations bonded to water as the hydronium ion,  $H_3O^+$ , read H- three O H, C-2) and conjugated anions/bases
- base substances which release hydroxide ions:  $OH^-$  (base O H, C-2)

### Brønsted-Lowry definitions:

- acid substances that donate a proton:  $H^+$ , also accepted
- base substances which accept a proton:  $H^+$ , also donated

Table 1: Representative reactions for  $H^+$  (also known as the Lewis acid-base reactions)

| Reaction                                                             | $K_c$                 |
|----------------------------------------------------------------------|-----------------------|
| 1. $HCl(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + Cl^-(aq)$      | $1.0 \times 10^6$     |
| 2. $HNO_3 + H_2O \rightleftharpoons H_3O^+ + NO_3^-$ (strong acid)   | $4.0 \times 10^9$     |
| 3. $HNO_2(aq) + H_2O \rightleftharpoons H_3O^+ + NO_2^-$ (weak acid) | $4.5 \times 10^{-4}$  |
| 4. $H_2SO_4 + H_2O \rightleftharpoons H_3O^+ + HSO_4^-$              | $1.0 \times 10^3$     |
| 5. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$                    | $1.0 \times 10^{-14}$ |

## Critical Thinking Questions

1. List acid-base conjugates for Table 1 that are Arrhenius acids:  
 $HCl(aq), HNO_3(aq), HNO_2(aq)$
2. List acid-base conjugates for Table 1 that are Arrhenius bases:  
 $Cl^-(aq), NO_3^-(aq)$
3. List acid-base conjugates in Table 1 that are Arrhenius strong acids:  
 $HCl(aq), HNO_3(aq), H_2SO_4(aq), H_2SO_4(aq), HNO_2(aq)$
4. List acid-base conjugates in Table 1 that are Arrhenius weak acids:  
 $HNO_2(aq), H_2O(l), H_2O(l)$
5. Is it possible for an acid to act as a base? Explain:  
 Yes, because  $H_2O(l), H_2O(l)$  can act as a base
6. Do acid substances contain Arrhenius-Brønsted bases within different molecules? Explain:  
 yes, because  $H_2O(l)$  is  $H_2O(l)$  and  $H_2O(l)$  can act as a base within  $H_2O(l)$
7. Is there any strong base that acts as a Brønsted-Lowry acid? Explain:

# Intro To Acids And Bases Worksheet Answers

## Chemquest 48

**Roman Wölfel**



## **Intro To Acids And Bases Worksheet Answers Chemquest 48:**

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The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts.

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