

BIO 110 Notes

8/21 Lecture 1

- Eukaryotic cells > prokaryotic cells > viruses > macromolecules > single atoms
- Atoms = building block of life
- Atomic Structure
 - Elements are made of a central nucleus that contains most of the atomic mass
 - The nucleus contains protons and neutrons, and is surrounded by electrons that orbit around the nucleus
 - Protons = +1 charge
 - Electrons = -1 charge
 - Neutrons = neutral
- Electrically neutral atoms: protons = electrons
- Imbalance of electrons = ions
 - Ions - net electric charge from either loss or gain of electrons
 - Forms ionic bonds - transfer of electrons
- Covalent bonds - sharing of electrons
 - Nonpolar - equal sharing of electrons
 - Polar - unequal sharing of electrons
- Electronegativity determines if something is polar or nonpolar
 - OR how strongly an atom attracts an electron
 - If EN is similar = nonpolar
 - If EN is different = polar
 - How different? If EN > or equal to 0.5 = polar covalent bond

8/26 Lecture 2 - Properties of water, acids and bases, macromolecules

- Importance of water
 - No life without water bc:
 - 1. Many biochemical reactions require water
 - 2. Important means of transport
 - 3. Very resistant to temperature change and acts as a buffer
- water is polar
 - held together by covalent bonds
 - oxygen is more electronegative than hydrogen, so the partial charges result in hydrogen bonds
 - Hydrogen bonds - attraction between hydrogen (partial positive) and a partial negative atom
- Polarity and solubility
 - Solubility - dissolves in water
 - Directly affected by polarity of water
 - Polar=soluble=hydrophilic (water loving)
 - Nonpolar = insoluble=hydrophobic (water fearing)
- Acids and Bases

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