



Nov 9, 2015, 9:41 AM



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Date: 11/9/15

Class: 9

5 INTERPRETING GRAPHICS

Use with Section 5.3



Figure 1 The emission spectrum and orbit-transition diagram for hydrogen.

speed of light
 $c = 3 \times 10^8 \text{ m/s}$
 $c = \lambda \cdot \nu$ *wave length*
 $E = h \cdot \nu$ *energy*
 $h = 6.63 \times 10^{-34} \text{ Js}$
Planck's constant

1) $2.46 \times 10^{-18} = 6.63 \times 10^{-34} \nu$
 $\nu = \nu$
 $\frac{2.46 \times 10^{-18}}{6.63 \times 10^{-34}} = 3.71 \times 10^{15} \text{ (Hz)}$
 $\frac{c}{\nu} = \lambda$

Transition	Energy, J (J)	ν (s ⁻¹)	λ (m)	Type of Radiation
$n=5 \rightarrow n=3$	2.66×10^{-19}	4.01×10^{15}	7.48×10^{-8}	IR
$n=5 \rightarrow n=4$	7.57×10^{-20}	1.14×10^{15}	2.63×10^{-7}	IR
$n=5 \rightarrow n=3$	1.82×10^{-19}	1.4×10^{15}	2.13×10^{-7}	IR
$n=5 \rightarrow n=2$	4.84×10^{-19}	6.76×10^{15}	4.44×10^{-8}	visible
$n=5 \rightarrow n=1$	2.82×10^{-18}	3.20×10^{15}	9.36×10^{-8}	UV
$n=5 \rightarrow n=4$	4.90×10^{-20}	7.40×10^{14}	4.01×10^{-7}	IR
$n=5 \rightarrow n=3$	1.55×10^{-19}	2.35×10^{15}	1.28×10^{-7}	IR
$n=5 \rightarrow n=2$	4.56×10^{-19}	6.91×10^{15}	4.34×10^{-8}	visible
$n=5 \rightarrow n=1$	2.69×10^{-18}	3.17×10^{15}	9.46×10^{-8}	UV
$n=4 \rightarrow n=3$	1.06×10^{-19}	1.61×10^{15}	1.86×10^{-7}	IR
$n=4 \rightarrow n=2$	4.09×10^{-19}	6.17×10^{15}	4.99×10^{-8}	visible
$n=4 \rightarrow n=1$	2.04×10^{-18}	3.08×10^{15}	9.74×10^{-8}	UV
$n=3 \rightarrow n=2$	3.03×10^{-19}	4.57×10^{15}	6.56×10^{-8}	visible
$n=3 \rightarrow n=1$	1.94×10^{-18}	2.93×10^{15}	1.21×10^{-7}	UV
$n=2 \rightarrow n=1$	1.64×10^{-18}	2.47×10^{15}	1.21×10^{-7}	UV

$h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$
 Planck's constant

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Lei Shi

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