



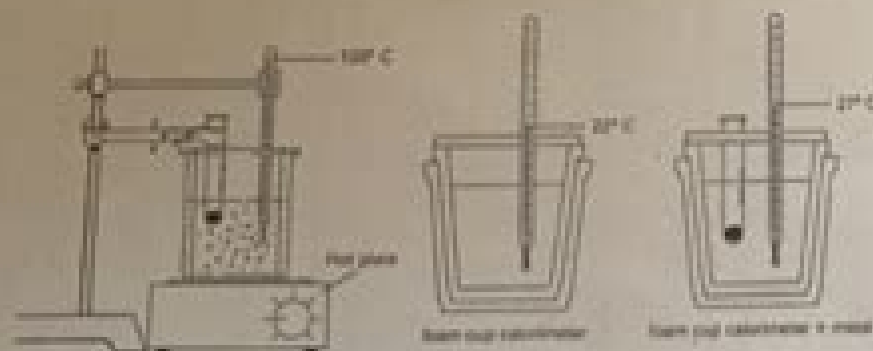
INTERPRETING GRAPHICS

USE WITH SECTION 11.1

A student performed an experiment to determine the specific heat of an unknown metal. The data she collected is reproduced in the table below. Use this information to answer the following questions.

Quantity	Trial 1	Trial 2
1. Mass of test tube + metal	118.19 g	118.21 g
2. Mass of test tube	47.88 g	47.88 g
3. Mass of metal	70.31 g	70.33 g
4. Mass of calorimeter	1.237 g	1.218 g
5. Mass of calorimeter + water	46.127 g	45.770 g
6. Mass of water	38.100 g	38.412 g
7. Initial temperature of metal	100.0 °C	100.0 °C
8. Initial temperature of water	22.8 °C	21.6 °C
9. Final temperature of water	27.8 °C	26.1 °C

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1. What was the final temperature of the metal?

a. Trial 1

b. Trial 2

2. What was the ΔT for the metal?

a. Trial 1

b. Trial 2

3. What was the ΔT for the water?

a. Trial 1

b. Trial 2

4. Calculate the heat change for the water. (Specific heat capacity for water is $4.184 \text{ J/g}^\circ\text{C}$)

a. Trial 1

b. Trial 2

5. Calculate the heat change for the metal.

a. Trial 1

b. Trial 2

6. Calculate the specific heat of the metal.

a. Trial 1

b. Trial 2

7. What metal might this have been?

a. aluminum

b. silver

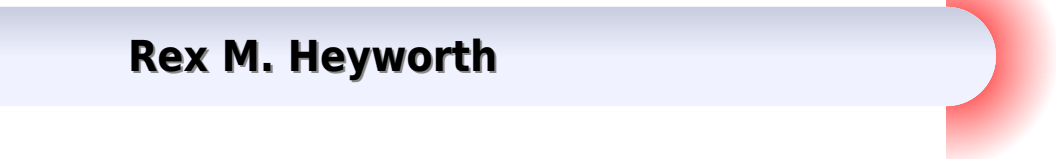
c. iron

d. mercury

88 Review Module 1 (Chapters 8-12)

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Rex M. Heyworth



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1. Understanding the eBook Interpreting Graphics Chemistry Pearson Answer Key
 - The Rise of Digital Reading Interpreting Graphics Chemistry Pearson Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Interpreting Graphics Chemistry Pearson Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Interpreting Graphics Chemistry Pearson Answer Key
 - User-Friendly Interface

4. Exploring eBook Recommendations from Interpreting Graphics Chemistry Pearson Answer Key
 - Personalized Recommendations
 - Interpreting Graphics Chemistry Pearson Answer Key User Reviews and Ratings
 - Interpreting Graphics Chemistry Pearson Answer Key and Bestseller Lists
5. Accessing Interpreting Graphics Chemistry Pearson Answer Key Free and Paid eBooks
 - Interpreting Graphics Chemistry Pearson Answer Key Public Domain eBooks
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 - ePub, PDF, MOBI, and More
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 - Interpreting Graphics Chemistry Pearson Answer Key Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Interpreting Graphics Chemistry Pearson Answer Key
 - Highlighting and Note-Taking Interpreting Graphics Chemistry Pearson Answer Key
 - Interactive Elements Interpreting Graphics Chemistry Pearson Answer Key
8. Staying Engaged with Interpreting Graphics Chemistry Pearson Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Interpreting Graphics Chemistry Pearson Answer Key
9. Balancing eBooks and Physical Books Interpreting Graphics Chemistry Pearson Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Interpreting Graphics Chemistry Pearson Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Interpreting Graphics Chemistry Pearson Answer Key
 - Setting Reading Goals Interpreting Graphics Chemistry Pearson Answer Key
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Interpreting Graphics Chemistry Pearson Answer Key
 - Fact-Checking eBook Content of Interpreting Graphics Chemistry Pearson Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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