

Part D Questions and Problems

26. An increase in the oxidation number of an atom indicates oxidation. A decrease in the oxidation number indicates reduction.
27. N is reduced (+5 to +2);
Br is oxidized (-1 to 0)
Mn is reduced (+7 to +2);
Cl is oxidized (-1 to 0)
N is reduced (+5 to +2);
Sb is oxidized (0 to +5)
S is reduced (+6 to +4);
C is oxidized (0 to +4)

Section Review 20.3

Part A Completion

- | | |
|---------------------|----------|
| 1. oxidation number | 5. two |
| 2. half-reaction | 6. added |
| 3. balanced | 7. ionic |
| 4. ionic | |

Part B True-False

- | | | |
|-------|--------|--------|
| 8. AT | 10. AT | 12. NT |
| 9. AT | 11. NT | 13. NT |

Part C Matching

- | | | |
|-------|-------|-------|
| 14. c | 17. b | 20. g |
| 15. a | 18. e | |
| 16. f | 19. d | |

Part D Questions and Problems

21. a. $\text{HNO}_3 \rightarrow \text{NO}$; N changes +5 to +2, a gain of $3e^-$; multiply by 2
 $2\text{HI} \rightarrow \text{I}_2$; I changes from -1 to 0, a loss of $2e^-$ for I_2 ; multiply by 3
 $2\text{HNO}_3 + 6\text{HI} \rightarrow 2\text{NO} + 3\text{I}_2 + 4\text{H}_2\text{O}$
- b. $\text{HNO}_3 \rightarrow \text{NO}_2$; N changes +5 to +4, a gain of $1e^-$; multiply by 10
 $\text{I}_2 \rightarrow 2\text{HIO}_3$; I changes from 0 to +5, a loss of $10e^-$ for I_2
 $10\text{HNO}_3 + \text{I}_2 \rightarrow 2\text{HIO}_3 + 10\text{NO}_2 + 4\text{H}_2\text{O}$
22. a. $\text{S}^{2-} \rightarrow \text{S} + 2e^-$ and
 $3e^- + 4\text{H}^+ + \text{NO}_3^- \rightarrow \text{NO} + 2\text{H}_2\text{O}$
Multiply the oxidation reaction by 3 and the reduction reaction by 2.
 $3\text{H}_2\text{S} + 2\text{HNO}_3 \rightarrow 3\text{S} + 2\text{NO} + 4\text{H}_2\text{O}$

- b. $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$ and $6e^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$; multiply the oxidation reaction by 6
 $14\text{H}^+ + 6\text{Fe}^{2+} + \text{Cr}_2\text{O}_7^{2-} \rightarrow 6\text{Fe}^{3+} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$

Practice Problems

Section 20.1

- | | |
|------------------------|---------------------------|
| 1. Sr: | oxidized (reducing agent) |
| O_2 : | reduced (oxidizing agent) |
| 2. Li: | oxidized (reducing agent) |
| S: | reduced (oxidizing agent) |
| 3. Cs: | oxidized (reducing agent) |
| Br_2 : | reduced (oxidizing agent) |
| 4. Mg: | oxidized (reducing agent) |
| N_2 : | reduced (oxidizing agent) |
| 5. Fe: | oxidized (reducing agent) |
| O_2 : | reduced (oxidizing agent) |
| 6. Br^- : | oxidized (reducing agent) |
| Cl_2 : | reduced (oxidizing agent) |
| 7. Se: | oxidized (reducing agent) |
| F_2 : | reduced (oxidizing agent) |
| 8. Ca: | oxidized (reducing agent) |
| O_2 : | reduced (oxidizing agent) |
| 9. Mg: | oxidized (reducing agent) |
| H^+ : | reduced (oxidizing agent) |
| 10. Na: | oxidized (reducing agent) |
| H_2O : | reduced (oxidizing agent) |

Section 20.2

- a. Sn is tin in an uncombined state. The oxidation number is 0.
- b. The ionic charge on potassium is $1+$, thus the oxidation number is +1.
- c. The ionic charge on sulfur is $2-$, thus the oxidation number is -2.
- d. The ionic charge on iron is $3+$, thus the oxidation number is +3.
- e. Se is selenium in an uncombined state. The oxidation number is 0.
- f. The ionic charge on magnesium is $2+$, thus the oxidation number is +2.
- g. The ionic charge on tin is $4+$, thus the oxidation number is +4.

Interpreting Graphics Answers Section Review

Malihe Alikhani



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