

# 1st Semester Chemistry Final Exam Review

1. Identify the following reaction type AND correct products:



- a. combustion
- b. synthesis
- c. decomposition
- d. single replacement
- e. double replacement

- w.  $\text{Ag} + \text{ScOH}$
- x.  $\text{Ag} + 2 \text{Sc}(\text{OH})_2$
- y.  $3 \text{Ag} + \text{Sc}(\text{OH})_3$
- z.  $3 \text{Ag} + 2 \text{Sc}(\text{OH})_4$

- |       |         |
|-------|---------|
| I.    | a. & w. |
| II.   | b. & z. |
| III.  | b. & y. |
| IV.   | d. & z. |
| V.    | e. & z. |
| VI.   | d. & y. |
| VII.  | e. & y. |
| VIII. | c. & z. |
| IX.   | d. & x. |

2. Convert:  $2.32 \times 10^5 \text{ mg} = ? \text{ kg}$

3. Calculate the energy for a quantum of light with a frequency of  $8.44 \times 10^{16} \text{ Hz}$ .

4. Determine the oxidation number for iodine in aluminum iodate.

5. REPORTING WITH THE CORRECT NUMBER OF SIGNIFICANT DIGITS, how many grams are in **4.011 moles** of calcium chloride?

- a) 445.14078 grams    b) 445.140 grams    c) 445.1 grams    d) 445 grams  
e) 450 grams    f) 400 grams

6. Calculate the frequency of light that has a wavelength of 210 nm.

# Honors Chemistry Practice Test

**RJ Alexander**



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