

Name Answer Key

Hour \_\_\_\_\_

## Chapter 4 Review

(Honors Chemistry)

Use the given number of neutrons and electrons to fill in the chart and determine whether the atom is...

Common: all information is exactly like the information on the periodic table.

Ion: the electron number differs from the proton number, but all other information is the same.

Isotope: the mass of the atom is different than the most common version.

	ELEMENT	ATOMIC NUMBER	MASS	PROTON NUMBER	NEUTRON NUMBER	ELECTRON NUMBER	ISOTOPE, ION OR COMMON
1.	<u>Sc</u>	21	45	<u>21</u>	<u>24</u>	18	<u>Cation, <math>Sc^{3+}</math></u>
2.	<u>K</u>	19	<u>42</u>	<u>19</u>	23	19	<u>Isotope</u>
3.	Cobalt	<u>27</u>	<u>59</u>	<u>27</u>	32	29	<u>Anion, <math>Co^{2-}</math></u>
4.	<u>S</u>	<u>16</u>	<u>32</u>	16	16	18	<u>Anion, <math>S^{2-}</math></u>
5.	Argon	<u>18</u>	40	<u>18</u>	<u>22</u>	18	<u>Common</u>
6.	Uranium	<u>92</u>	<u>238</u>	<u>92</u>	146	92	<u>Common</u>
7.	<u>Ca</u>	<u>20</u>	<u>43</u>	20	23	20	<u>Isotope</u>
8.	<u>P</u>	15	<u>31</u>	<u>15</u>	16	18	<u>Anion, <math>P^{3-}</math></u>
9.	<u>Pb</u>	<u>82</u>	<u>225</u>	82	143	82	<u>Isotope</u>
10.	Zinc	<u>30</u>	<u>65</u>	<u>30</u>	35	30	<u>Common</u>
11.	Aluminum	<u>13</u>	30	<u>13</u>	<u>17</u>	13	<u>Isotope</u>
12.	Potassium	<u>19</u>	<u>39</u>	<u>19</u>	20	18	<u>Cation, <math>K^{+}</math></u>
13.	<u>Ce</u>	58	<u>142</u>	<u>58</u>	84	58	<u>Isotope</u>
14.	Sodium	<u>11</u>	23	<u>11</u>	<u>12</u>	10	<u>Cation, <math>Na^{+}</math></u>
15.	Chlorine	<u>17</u>	<u>35</u>	<u>17</u>	18	18	<u>Anion, <math>Cl^{-}</math></u>
16.	Radon	<u>86</u>	222	<u>86</u>	<u>136</u>	86	<u>Common</u>
17.	<u>Mg</u>	12	22	<u>12</u>	<u>10</u>	12	<u>Isotope</u>
18.	Boron	<u>5</u>	<u>11</u>	<u>5</u>	6	5	<u>Common</u>
19.	Nickel	<u>28</u>	59	<u>28</u>	<u>31</u>	26	<u>Cation, <math>Ni^{2+}</math></u>
20.	Bismuth	<u>83</u>	209	<u>83</u>	<u>126</u>	83	<u>Common</u>
21.	Iridium	<u>77</u>	<u>196</u>	<u>77</u>	119	77	<u>Isotope</u>
22.	<u>Hg</u>	80	<u>201</u>	<u>80</u>	121	78	<u>Cation, <math>Hg^{2+}</math></u>
23.	Titanium	22	50	<u>22</u>	<u>28</u>	22	<u>ISOTOPE</u>

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**Emilie Sanchez**



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