

## GAS STOICHIOMETRY WORKSHEET

Please answer the following on separate paper using proper units and showing all work. Please note that these problems require a **balanced chemical equation**.

- Carbon monoxide reacts with oxygen to produce carbon dioxide. If 1.0 L of carbon monoxide reacts with oxygen at STP,
  - how many liters of oxygen are required to react?
  - How many liters of carbon dioxide are produced?
- Acetylene gas ( $C_2H_2$ ) undergoes combustion to produce carbon dioxide and water vapor.
  - How many liters of  $C_2H_2$  are required to produce 75.0 L of  $CO_2$ ?
  - What volume of  $H_2O$  is produced?
  - What volume of  $O_2$  is required?
- If liquid carbon disulfide ( $CS_2$ ) reacts with 450 mL of oxygen to produce the gases carbon dioxide and sulfur dioxide, what volume of each product is produced?
- Assume that 5.60 L of hydrogen gas at STP reacts with copper (II) oxide according to the following balanced equation:
$$CuO (s) + H_2 (g) \rightarrow Cu (s) + H_2O (g)$$
  - How many moles of  $H_2$  react?
  - How many moles of copper are produced?
  - How many grams of copper are produced?
- Assume that 8.5 L of iodine gas ( $I_2$ ) are produced at STP according to the following balanced equation:
$$2 KI (aq) + Cl_2 (g) \rightarrow 2 KCl (aq) + I_2 (g)$$
  - How many moles of  $I_2$  are produced?
  - How many moles of KI were used?
  - How many grams of KI were used?
- Solid iron (III) hydroxide decomposes to produce iron (III) oxide and water vapor. If 0.75 L of water vapor are produced at STP,
  - How many grams of iron (III) hydroxide were used?
  - How many grams of iron (III) oxide were produced?

# Gas Stoichiometry Practice Sheet 20cavalcade

**Marcel A. Müller**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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