

Questions and Problems

Q1 According to your graph, what is the relationship between the pressure and the volume of a gas?

* As the volume of the gas increases the pressure of the gas decreases. They have inverse proportional relationship.

Q2 On your graph, what is the volume of the gas at a pressure of 760. mmHg?

* After plotting the data to excel, the graph produced a linear equation $y = -28.873x + 1539.3$ where y is the pressure and x is the volume of gas.

We need to find x at 760 mmHg.

Hence,

$$760 \text{ mmHg} = -28.873x + 1539.8 \Rightarrow x = 26.99 \text{ mL}$$

Therefore, at 760 mmHg the volume of the gas is 26.99 mL

Q3 On your graph, what is the pressure of the gas when the volume is 30.0 mL?

* $x = 30.0 \text{ mL}$, $y = ?$

$$y = -28.873 (30.0) + 1539.3$$

$$y = 673.1 \text{ mmHg}$$

Therefore, when the volume of the gas is 30.0 mL the pressure is 673.1 mmHg

Q4 A sample of helium has a volume of 325 mL and a pressure of 655 mmHg. What will be the pressure, in mmHg, if the sample of helium is compressed to 125 mL. (T, n constant)?

$$P_1 V_1 = P_2 V_2$$

$$P_2 = (P_1 V_1) \div V_2 = (655 \text{ mmHg} \times 325 \text{ mL}) \div 125 \text{ mL} = 1,703 \text{ mmHg}$$

Gas Laws Section 2 Review Answers

Anup Goel



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