

Skill - Student will accurately describe the relationship between P, V, T and n of gases.

1. if the temperature increases, the volume will...

INCREASE, an inc in temp causes an inc in volume due to a direct relationship.

2. if the pressure decreases the volume will...

INCREASE, a dec in pressure will cause an inc in volume due to an inverse relationship.

3. if the moles of gas decrease the volume will...

DECREASE, a dec in moles will cause a dec in volume due to a direct relationship.

Skill - Student will accurately calculate a 2-variable gas law problem.

4. A sample of methane gas that has a volume of 10.0 L at 25°C is cooled to 5°C at a constant pressure. Calculate the new volume. What equation must you use? **SHOW YOUR WORK.**

$$\frac{V_2}{T_2} = \frac{V_1}{T_1} \rightarrow \frac{V_2}{273K} = \frac{10.0L}{298K} \rightarrow V_2 = 9.36L$$

5. A gas occupies 12.3 liters at a pressure of 40.0 mmHg. What is the volume when the pressure is increased to 60.0 mmHg?

$$P_1V_1 = P_2V_2 \rightarrow (40 \text{ mmHg})(12.3L) = (60 \text{ mmHg})(V_2) \rightarrow V_2 = 8.2L$$

Skill - Student will accurately calculate using the ideal gas law.

6. Radon, a radioactive gas formed naturally in the soil, can cause lung cancer. It can pose a hazard to humans by seeping in to houses and become trapped in basements. A 1.5 mol sample of radon has a volume of 23.0 L at 33°C. What is the pressure of the gas? What equation must you use? **SHOW YOUR WORK.**

$$PV = nRT \rightarrow P = nRT/V \\ = (1.5 \text{ mol} \times 0.08206L \cdot \text{atm/Kmol} \times 306K) / 23.0L \rightarrow P = 1.8 \text{ atm}$$

7. A 3.25g sample of solid calcium carbide (CaC_2) reacts with water to produce acetylene gas (C_2H_2) and aqueous calcium hydroxide. If the temperature and pressure within the system are 17°C and 740.0 mmHg, how many milliliters of acetylene were produced? **SHOW YOUR WORK.**



$$3.25g \text{ CaC}_2 \times (1 \text{ mol CaC}_2 / 64.1g \text{ CaC}_2) \times (1 \text{ mol C}_2\text{H}_2 / 1 \text{ mol CaC}_2) = 0.051 \text{ mol C}_2\text{H}_2$$

$$PV = nRT \rightarrow V = nRT/P = (0.051 \text{ mol} \times 0.08206L \cdot \text{atm/Kmol} \times 290K) / 0.978 \text{ atm} = 1.24L = 1240 \text{ mL}$$

Gas Laws Review Answer Key

JA Banks



Gas Laws Review Answer Key:

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