

7. If you have 17 liters of gas at a temperature of 67 °C and a pressure of 88.89 atm, what will be the pressure of the gas if you raise the temperature to 94 °C and decrease the volume to 12 liters?

$$P_2 = \frac{P_1 V_1 T_2}{V_2 T_1} = \frac{88.89 \text{ atm} (17 \text{ L}) 367 \text{ K}}{(12 \text{ L}) 340 \text{ K}}$$

$$P_2 = \quad \text{atm}$$

8. I have an unknown volume of gas at a pressure of 0.5 atm and a temperature of 325 K. If I raise the pressure to 1.2 atm, decrease the temperature to 320 K, and measure the final volume to be 48 liters, what was the initial volume of the gas?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{1.2 \text{ atm} (48 \text{ L}) 325 \text{ K}}{320 \text{ K} (0.5 \text{ atm})}$$

$$V_1 = \quad \text{L}$$

9. If you have 21 liters of gas held at a pressure of 78 atm and a temperature of 900 K, what will be the volume of gas if you decrease the pressure to 45 atm and decrease the temperature to 750 K?

$$V_2 = \frac{78 \text{ atm} (21 \text{ L}) 750 \text{ K}}{900 \text{ K} (45 \text{ atm})}$$

$$V_2 = \quad \text{L}$$

10. If I have 2.9 L of gas at a pressure of 5 atm and a temperature of 50 °C, what will the temperature of the gas be if I decrease the volume of the gas to 2.4 L and decrease the pressure to 3 atm?

$$1/T_2 = \frac{P_1 V_1}{T_1 P_2 V_2} = \frac{(2.9 \text{ L}) 5 \text{ atm}}{323 \text{ K} (3 \text{ atm}) 2.4 \text{ L}}$$

$$T_2 = 160 \text{ K}$$

11. You have an unknown volume of gas held at a temperature of 115 K in a container with a pressure of 60 atm. If by increasing the temperature to 225 K and decreasing the pressure to 30 atm causes the volume of the gas to be 29 L, how many liters of gas did you start with?

$$V_1 = \frac{P_2 V_2 T_1}{T_2 P_1} = \frac{115 \text{ K} (29 \text{ L}) 30 \text{ atm}}{(225 \text{ K}) 60 \text{ atm}}$$

$$V_1 = \quad \text{L}$$

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S. K. Roy



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

Thermodynamics Problem Solving in Physical Chemistry Kathleen E. Murphy, 2020-03-23 Thermodynamics Problem Solving in Physical Chemistry Study Guide and Map is an innovative and unique workbook that guides physical chemistry students through the decision making process to assess a problem situation, create appropriate solutions, and gain confidence through practice solving physical chemistry problems. The workbook includes six major sections with 20-30 solved problems in each section that span from easy single objective questions to difficult multistep analysis problems. Each section of the workbook contains key points that highlight major features of the topic to remind students of what they need to apply to solve problems in the topic area. Key Features: Provides instructor access to a visual map depicting how all equations used in thermodynamics are connected and how they are derived from the three major energy laws. Acts as a guide in deriving the correct solution to a problem. Illustrates the questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry. Can be used as a stand alone product for review of Thermodynamics questions for major tests.

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Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance

Dr. Salm's Chemistry Problem Drill Book Robert J. Salm, Robert J. Salm Robert J. Salm Ph.D., 2005 This drill book contains many common problem types that are asked in General Chemistry classes in High School and College This work will give you practice with the major problem types as you prepare for finals and standardized tests

Thermofluids Michael Horsley, Sherwin Keith, 1996-12-07 The two associated subjects of thermodynamics and fluid mechanics are combined in this book to provide the reader with an easy to follow text which emphasizes the essential coherence of the material

Chemical Calculations with Explanatory Notes, Problems, and Answers, Specially Adapted for Use in Colleges and Science Schools Richard Lloyd Whiteley, 1896

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