

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Introduction

Standard Conditions

Unit	Standard Pressure
atm	1.013
mmHg	760
cmHg	76
torr	760

22.4 L = 1 mol at STP

• STP (standard temperature and pressure) is 0°C and 1 atm.

• The gas laws:

Gas Law	Equation	Variables
Boyle's Law	$P_1V_1 = P_2V_2$	P, V
Charles's Law	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$	V, T
Gay-Lussac's Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	P, T
Avogadro's Law	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$	V, n
Combined Gas Law	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$	P, V, T
Ideal Gas Law	$PV = nRT$	P, V, n, T

• The constant R used in the ideal gas law depends on the unit of pressure.

a. Units: $R = 0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}}$

b. Units: $R = 62.36 \frac{\text{torr} \cdot \text{L}}{\text{mol} \cdot \text{K}}$

Notes:

- Temperature must always be in Kelvin: $K = ^\circ\text{C} + 273$.
- If you are given two different units of pressure, convert first.
- Create two columns below (initial and final). Write what you're given in the problem. Keep the two that stay the same (variables you're given, and solve for the unknown). For ideal gas laws, write what you are given, insert those in equation and solve for unknown.
- Avogadro's Law: Solve for volume to use ideal standard values.
- Graham's Law: You will usually have the molecular mass for these problems. Double and triple check you have the right of effusion and molecular masses in the correct spots before solving the equation (this is an inverse relationship). You will calculate the square root or square both sides of the equation before moving things.

Charles' Law $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

1. The temperature of a gas is directly proportional to its volume. If the temperature of a gas is increased, the volume will also increase. If the temperature of a gas is decreased, the volume will also decrease.

2. A gas at 27°C and 1.0 atm has a volume of 2.0 L. What is the volume of the gas at 127°C and 1.0 atm?

Boyle's Law $P_1V_1 = P_2V_2$

1. The pressure of a gas is inversely proportional to its volume. If the pressure of a gas is increased, the volume will decrease. If the pressure of a gas is decreased, the volume will increase.

2. A gas at 1.0 atm and 2.0 L has a volume of 2.0 L. What is the pressure of the gas at 1.0 atm and 1.0 L?

Gay-Lussac's Law $\frac{P_1}{T_1} = \frac{P_2}{T_2}$

1. The pressure of a gas is directly proportional to its temperature. If the pressure of a gas is increased, the temperature will also increase. If the pressure of a gas is decreased, the temperature will also decrease.

2. A gas at 1.0 atm and 27°C has a pressure of 1.0 atm. What is the pressure of the gas at 127°C and 1.0 atm?

Avogadro's Law $\frac{V_1}{n_1} = \frac{V_2}{n_2}$

1. The volume of a gas is directly proportional to the number of moles of gas. If the volume of a gas is increased, the number of moles will also increase. If the volume of a gas is decreased, the number of moles will also decrease.

2. A gas at 1.0 atm and 2.0 L has a volume of 2.0 L. What is the volume of the gas at 1.0 atm and 1.0 L?

Combined Gas Law $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

1. The combined gas law is a combination of Boyle's Law, Charles's Law, and Gay-Lussac's Law. It relates the pressure, volume, and temperature of a gas.

2. A gas at 1.0 atm, 2.0 L, and 27°C has a pressure of 1.0 atm, a volume of 2.0 L, and a temperature of 27°C. What is the pressure, volume, and temperature of the gas at 1.0 atm, 1.0 L, and 127°C?

Ideal Gas Law $PV = nRT$

1. The ideal gas law is a combination of Boyle's Law, Charles's Law, Gay-Lussac's Law, and Avogadro's Law. It relates the pressure, volume, number of moles, and temperature of a gas.

2. A gas at 1.0 atm, 2.0 L, and 27°C has a pressure of 1.0 atm, a volume of 2.0 L, and a temperature of 27°C. What is the pressure, volume, number of moles, and temperature of the gas at 1.0 atm, 1.0 L, and 127°C?

All Gas Laws Included!



Gas Laws Packet Answers

A Gutmann



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Proposed Changes to Natural Gas Laws United States. Congress. House. Committee on Energy and Commerce. Subcommittee on Fossil and Synthetic Fuels, 1983

The Science Teacher's Activity-A-Day, Grades 5-10 Pam Walker, Elaine Wood, 2010-10-05 A hands on and fun filled resource for teaching science to middle and high school students New in the 5 Minute Fundamentals Series The Science Teacher s Activity A Day Grades 6 12 includes 180 easy five minute hook or sponge activities to capture learners attention and introduce lessons Divided into three units Physical Science Life Science and Earth and Space Science the activities cover topics based on the National Science Education Standards All the book s activities can be done with materials that are inexpensive and easy to find Includes quick and fun sponge activities that are designed to engage students All the activities take about 5 minutes to complete The Science Teacher s Activity a Day is an ideal resource for middle and high school science teachers

The Handy Chemistry Answer Book Justin P. Lomont, Ian C. Stewart, 2013-10-01 Don t be mixed up about chemistry Simplify the complex chemical reactions that take place everywhere in our lives with this engaging easy to follow question and answer guide Where would we be without atoms and compounds Gas liquids solids and plasma Acids and bases Bonds and reactions Matter and energy The Handy Chemistry Answer Book covers the building blocks of life and the universe The secret life of atoms how polar bears aren t actually white why oil and water don t mix and much much more are revealed and explained This informative guide covers the basics of chemistry history atomic structures chemical bonds and reactions organic and inorganic chemistry to more advanced material nuclear chemistry biochemistry physical and theoretical chemistry by answering nearly 1 000 common chemistry questions including What causes lightning How does photosynthesis work What are hard and soft Lewis acids and bases What makes a fabric waterproof What are the twelve principles of green chemistry When did alchemists finally abandon trying to make gold What is Le Chatelier s principle What do the different octane ratings mean at the gas pump What is genetic engineering Why is calcium important for strong bones What is the 18 electron rule Why does chocolate turn white as it ages Chemical reactions that rule the world their properties structure composition behavior and history are tackled and explained in plain English in The Handy Chemistry Answer Book With many photos illustrations a few formulas molecular diagrams and other graphics this fun fact filled tome is richly illustrated A history of chemistry timeline appendices on Nobel Prize in Chemistry winners a bibliography further reading section glossary of terms a table of physical constants a table of conversion factors and extensive index add to its usefulness

Atmospheric Thermodynamics Gerald R. North, Tatiana L. Erukhimova, 2009-04-09 This textbook presents a uniquely integrated approach in linking both physics and chemistry to the study of atmospheric thermodynamics The book explains the classical laws of thermodynamics focuses

on various fluid systems and recognising the increasing importance of chemistry in the meteorological and climate sciences devotes a chapter to chemical thermodynamics which includes an overview of photochemistry Although students are expected to have some background knowledge of calculus general chemistry and classical physics the book provides set aside refresher boxes as useful reminders It contains over 100 diagrams and graphs to supplement the discussions and a similar number of worked examples and exercises with solutions included at the end of the book It is ideal for a single semester advanced course on atmospheric thermodynamics and will prepare students for higher level synoptic and dynamics courses

Instructional Technology Research, Design and Development: Lessons from the Field Alias, Nor

Aziah, 2011-11-30 Design and development research which has considerable implications for instructional design focuses on designing and exploring products artifacts and models as well as programs activity and curricula Instructional Technology Research Design and Development Lessons from the Field is a practical text on design and development research in the field of instructional technology This book gives readers an overview of design and development research and how it is conducted in different contexts and for various purposes Further this reference source provides readers with practical knowledge on design and development research gained through investigation of lessons learned in the field

Learning and Leading with Technology, 1996

Applied Thermodynamics for Meteorologists Sam Miller, 2015-06-04 This textbook on atmospheric thermodynamics is for students of meteorology or atmospheric science It also serves as a reference text for working professionals in meteorology and weather forecasting It is unique because it provides complete calculus based derivations of basic physics from first principles and connects mathematical relationships to real world practical weather forecasting applications Worked examples and practice problems are included throughout

Oil and Gas Law: 1922-1951
Texas Law Review, 1954

Law Day U.S.A. Planning Guide and Program Manual American Bar Association, 1968

Water: Fundamentals as the Basis for Understanding the Environment and Promoting Technology P.G. Debenedetti, M.A. Ricci, F. Bruni, 2015-05-18 Water fascinating in all its phases forms and states of aggregation Without it life as we know it would not exist on Earth for as Paracelsus stated in the 16th century water is the matrix of the world and of all its creatures While it may appear to be a simple molecule there is still much about it which is not fully understood What is notably lacking is a microscopically based understanding of the reasons for the many anomalous properties of water This book presents lectures from the Enrico Fermi summer school Water fundamentals as the basis for understanding the environment and promoting technology held in Varenna Italy in July 2013 The aim of the school was to offer a glimpse of the many questions that remain unanswered about this molecule and topics covered included water in relation to other liquids biological water local environment of water protons atmospheric water amorphous solid phases of water NMR studies of water spectroscopic studies of water the structure of liquid water and supercooled water among others While this list is by no means exhaustive or complete it is wide enough to provide a solid basis to young researchers in the field and the book will be a valuable source

of reference for students and all those with an interest in the properties of this fascinating substance **United States Government Organization Manual** ,1969 **Terrestrial Hydrometeorology** W. James Shuttleworth,2012-01-30 Both hydrologists and meteorologists need to speak a common scientific language and this has given rise to the new scientific discipline of hydrometeorology which deals with the transfer of water and energy across the land atmosphere interface Terrestrial Hydrometeorology is the first graduate level text with sufficient breadth and depth to be used in hydrology departments to teach relevant aspects of meteorology and in meteorological departments to teach relevant aspects of hydrology and to serve as an introductory text to teach the emerging discipline of hydrometeorology The book will be essential reading for graduate students studying surface water hydrology meteorology and hydrometeorology It can also be used in advanced undergraduate courses and will be welcomed by academic and professional hydrologists and meteorologists worldwide Additional resources for this book can be found at <http://www.wiley.com/go/shuttleworth> hydrometeorology **Resources in Education** ,1993 **Texas Oil and Gas Law Journal** ,2002 *Meteorology* ,1971

A First Course in Atmospheric Thermodynamics Grant W. Petty,2008-01-01 This textbook is written for meteorology majors who require an initial introduction to the physical properties of the atmosphere and to the essential principles and real world applications of atmospheric thermodynamics These topics are supplemented by a sampling of techniques and technologies related to atmospheric measurements and observations A unique tutorial included as an appendix teaches students how to attack physical problems symbolically deferring numerical calculations until the final step in the solution The author s objectives include not only covering the traditional core subject matter of an undergraduate thermodynamics course but also facilitating students transition from a purely abstract understanding of calculus and physics concepts to the confident application of both to the science of meteorology **Examining Impacts of Federal Natural Resources Laws**

Gone Astray United States. Congress. House. Committee on Natural Resources. Subcommittee on Oversight and Investigations,2017 **Examining Impacts of Federal Natural Resources Laws Gone Astray, Part II, Serial No. 115-16, July 18, 2017, 115-1** ,2017 **Business Law** Dr. S. K. Singh, Dr. Angad Tiwary, Nikhil Gupta,2021-05-31 Unit I Indian Contract Act 1872 1 Business Mercantile Law An Introduction 2 Indian Contract Act 1872 An Introduction 3 Contract Meaning Definition and Characteristics of a Valid Contract 4 Agreement Meaning Kinds and Difference 5 Proposal Offer Acceptance Communication and Revocation 6 Capacity of Parties to Contract or Parties Competency to Contract 7 Free Consent 8 Lawful Consideration and Objects 9 Agreements Expressly Declared as Void 10 Contingent Contracts 11 Performance of Contracts and Appropriation of Payments 12 Discharge of Contracts 13 Quasi or Implied Contracts of Certain Relations Resembling those Created by Contracts Sections 68 to 72 14 Remedies for Breach of Contract Unit II Special Contracts 15 Contract of Indemnity and Guarantee 16 Contract of Bailment and Pledge 17 Contracts of Agency Unit III The Sale of Goods Act 1930 18 The Sale of Goods Act 1930 An Introduction 19 Conditions and Warranties 20 Effects of the

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