

**Gas Laws Worksheet**

$$\text{atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

**Boyle's Law Problems:**

$$P_1 V_1 = P_2 V_2$$

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(748 \text{ mmHg})(22.5 \text{ L}) = (725 \text{ mmHg}) V_2$$

$$V_2 = \frac{(748 \text{ mmHg})(22.5 \text{ L})}{(725 \text{ mmHg})}$$

$$V_2 = 23.2 \text{ L}$$

2. A gas with a volume of 4.0 L at a pressure of 205 kPa is allowed to expand to a volume of 12.0 L. What is the pressure in the container if the temperature remains constant?

$$(4.0 \text{ L})(205 \text{ kPa}) = (12.0 \text{ L}) P_2$$

$$P_2 = \frac{(4.0 \text{ L})(205 \text{ kPa})}{12.0 \text{ L}}$$

$$P_2 = 68.3 \text{ kPa}$$

3. What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

$$(196.0 \text{ L})(1.00 \text{ atm}) = (26.0 \text{ L}) P_2$$

$$P_2 = \frac{(196.0 \text{ L})(1.00 \text{ atm})}{26.0 \text{ L}}$$

$$P_2 = 7.54 \text{ atm}$$

4. A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

$$(40.0 \text{ L})(12.7 \text{ kPa}) = (8.4 \text{ kPa}) V_2$$

$$V_2 = \frac{(40.0 \text{ L})(12.7 \text{ kPa})}{8.4 \text{ kPa}}$$

$$V_2 = 60.5 \text{ L}$$

# Gas Law Problems Guided Practice

**Pamela Mills, Amina El-Ashmawy**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author names.

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**Survival Guide to General Chemistry** Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

**The Practice of Chemistry Study Guide & Solutions Manual** Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes. Also includes solutions and explanations to odd numbered exercises.

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## **Table of Contents Gas Law Problems Guided Practice**

1. Understanding the eBook Gas Law Problems Guided Practice
  - The Rise of Digital Reading Gas Law Problems Guided Practice
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Law Problems Guided Practice
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Law Problems Guided Practice
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Law Problems Guided Practice
  - Personalized Recommendations
  - Gas Law Problems Guided Practice User Reviews and Ratings

- Gas Law Problems Guided Practice and Bestseller Lists
- 5. Accessing Gas Law Problems Guided Practice Free and Paid eBooks
  - Gas Law Problems Guided Practice Public Domain eBooks
  - Gas Law Problems Guided Practice eBook Subscription Services
  - Gas Law Problems Guided Practice Budget-Friendly Options
- 6. Navigating Gas Law Problems Guided Practice eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Law Problems Guided Practice Compatibility with Devices
  - Gas Law Problems Guided Practice Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Law Problems Guided Practice
  - Highlighting and Note-Taking Gas Law Problems Guided Practice
  - Interactive Elements Gas Law Problems Guided Practice
- 8. Staying Engaged with Gas Law Problems Guided Practice
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Law Problems Guided Practice
- 9. Balancing eBooks and Physical Books Gas Law Problems Guided Practice
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Law Problems Guided Practice
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Law Problems Guided Practice
  - Setting Reading Goals Gas Law Problems Guided Practice
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Law Problems Guided Practice
  - Fact-Checking eBook Content of Gas Law Problems Guided Practice
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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