

Chapter Three ENTROPIC FLOW OF A PERFECT GAS

Problem 11. An isobaric piston-cylinder device is initially at 1 bar and a volume of 100 cm³. The temperature pressure of the gas is 100 K. The temperature dependence of the heat capacity of the gas is given by the chemical formula: $C_p = 10 + 0.01T$ J K⁻¹ mol⁻¹, where T is the temperature in Kelvin.

$$(a) \quad \int_{100}^{200} C_p dT = 10(200 - 100) + 0.01 \int_{100}^{200} T dT = 1000 + 0.5 J$$

$$(b) \quad \int_{100}^{200} C_p dT = 10(200 - 100) + 0.01 \int_{100}^{200} T dT = 1000 + 0.5 J$$

$$(c) \quad \frac{C_p}{R} = \frac{10}{8.314} + \frac{0.01T}{8.314} = 1.2016 + 0.0012028T$$

$$(d) \quad \int_{100}^{200} C_p dT = 1000 + 0.5 J$$

$$(e) \quad \int_{100}^{200} C_p dT = 1000 + 0.5 J$$

$$(f) \quad \int_{100}^{200} C_p dT = 1000 + 0.5 J$$

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Michael Rajnik



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