

M A buffer is made using the conjugate acid-base pair acetic acid and acetate ions. Write chemical equations showing the reactions that take place when H^+ and when OH^- are added to the buffer.

4. Acetic acid has a K_a value of 1.8×10^{-5} . How many grams of $\text{NaC}_2\text{H}_3\text{O}_2$ would have to be added to 100 ml. of 0.10 M $\text{HC}_2\text{H}_3\text{O}_2$ to prepare a buffer with a pH of 4.50?

$$(0.057M)(.100L) = 0.0057\text{mol}$$
$$= 0.057 M$$

Acetic acid, $\text{HC}_2\text{H}_3\text{O}_2$, 0.1 M, 20 ml.

Sodium chloride, NaCl, 0.1 M, 40 ml.

Ammonia, 50% 0.1 M, 20 ml.

Sodium Hydroxide, NaOH, 0.1 M, 5 ml.

Ammonium chloride, NH_4Cl , 0.1 M, 20 ml.

Sodium hydroxide, NaOH, 0.2 M, 30 ml.

Hydrochloric acid, HCl, 0.1 M, 5 ml.

Unknown solid acid anhydride, 2.0 g

Phenylhydrazine, 1.0 g, 1 ml

Water, distilled or deionized

Sodium acetate, $\text{NaC}_2\text{H}_3\text{O}_2$, 0.1 M, 20 ml.

$$[\text{C}_2\text{H}_3\text{O}_2^-] = \frac{K_a [\text{HC}_2\text{H}_3\text{O}_2]}{[\text{H}^+]} = \frac{(1.8 \times 10^{-5})(0.10)}{3.16 \times 10^{-5}}$$

Halterman, G. L., et al.

2

Petersburg, 1992—and, of

CLIMATE 195

Abstract—Cancer patients 3

Buret, ring stand, and bevel clamp

Wash bottle with distilled or deionized water.

Reference: 150 mL, 1

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M Mark



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