

## Masses and Moles

Look at this example of how to calculate the mass from the number of moles:

Mass of substance = number of moles  $\times$  mass of 1 mole

*Which is the same as saying....*

Mass = number of moles  $\times$  Relative Formula Mass

$$m = n \times M_r$$

### Example

What is the mass of 2 moles of carbon?

$m$	$=$	$n \times M_r$
	$=$	$2 \times 12$
	$=$	$24g$

Use the Periodic Table in the front of the book to answer these questions:

**Q1** What is the mass of...?

- |  |  |
|--|--|
| <ul style="list-style-type: none"> <li>a) 1 mole of carbon atoms</li> <li>b) 2 moles of sodium atoms</li> <li>c) 2 moles of aluminium atoms</li> <li>d) 3 moles of lithium atoms</li> <li>e) 10 moles of iodine atoms</li> </ul> | <ul style="list-style-type: none"> <li>f) 1 mole of oxygen molecules</li> <li>g) 2 moles of chlorine molecules</li> <li>h) 1 mole of nitrogen molecules</li> <li>i) 0.5 moles of nitrogen molecules</li> <li>j) 0.1 moles of carbon molecules</li> </ul> |
|--|--|

**Q2** What is the mass of...?

- |  |  |
|--|--|
| <ul style="list-style-type: none"> <li>a) 1 mole of carbon dioxide (<math>\text{CO}_2</math>)</li> <li>b) 5 moles of carbon dioxide (<math>\text{CO}_2</math>)</li> <li>c) 3 moles of water (<math>\text{H}_2\text{O}</math>)</li> <li>d) 2 moles of sulphur trioxide (<math>\text{SO}_3</math>)</li> <li>e) 3 moles of nitric acid (<math>\text{HNO}_3</math>)</li> </ul> | <ul style="list-style-type: none"> <li>f) 10 moles of hydrochloric acid (<math>\text{HCl}</math>)</li> <li>g) 100 moles of sulphuric acid (<math>\text{H}_2\text{SO}_4</math>)</li> <li>h) 50 moles of sodium hydroxide (<math>\text{NaOH}</math>)</li> <li>i) 30 moles of calcium oxide (<math>\text{CaO}</math>)</li> <li>j) 25 moles of sodium chloride (<math>\text{NaCl}</math>)</li> </ul> |
|--|--|

**Q3** What is the mass of...?

- |  |  |
|--|--|
| <ul style="list-style-type: none"> <li>a) 0.1 moles of copper oxide (<math>\text{CuO}</math>)</li> <li>b) 0.1 moles of calcium carbonate (<math>\text{CaCO}_3</math>)</li> <li>c) 0.01 moles of sulphuric acid (<math>\text{H}_2\text{SO}_4</math>)</li> <li>d) 0.2 moles of calcium chloride (<math>\text{CaCl}_2</math>)</li> <li>e) 0.25 moles of sodium carbonate (<math>\text{Na}_2\text{CO}_3</math>)</li> </ul> | <ul style="list-style-type: none"> <li>f) 0.05 moles of sodium hydroxide (<math>\text{NaOH}</math>)</li> <li>g) 0.58 moles of aluminium oxide (<math>\text{Al}_2\text{O}_3</math>)</li> <li>h) 0.1 moles of propanoic acid (<math>\text{C}_2\text{H}_5\text{COOH}</math>)</li> <li>i) 0.5 moles of potassium manganate(VII) (<math>\text{KMnO}_4</math>)</li> <li>j) 0.25 moles of Hydrated Magnesium Sulphate (<math>\text{MgSO}_4 \cdot 7\text{H}_2\text{O}</math>)</li> </ul> |
|--|--|

# Moles Chemistry Mole Questions And Answers

**S Ben Porath**



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## **Table of Contents Moles Chemistry Mole Questions And Answers**

1. Understanding the eBook Moles Chemistry Mole Questions And Answers
  - The Rise of Digital Reading Moles Chemistry Mole Questions And Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Moles Chemistry Mole Questions And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Moles Chemistry Mole Questions And Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Moles Chemistry Mole Questions And Answers
  - Personalized Recommendations
  - Moles Chemistry Mole Questions And Answers User Reviews and Ratings
  - Moles Chemistry Mole Questions And Answers and Bestseller Lists
5. Accessing Moles Chemistry Mole Questions And Answers Free and Paid eBooks
  - Moles Chemistry Mole Questions And Answers Public Domain eBooks
  - Moles Chemistry Mole Questions And Answers eBook Subscription Services
  - Moles Chemistry Mole Questions And Answers Budget-Friendly Options

6. Navigating Moles Chemistry Mole Questions And Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Moles Chemistry Mole Questions And Answers Compatibility with Devices
  - Moles Chemistry Mole Questions And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Moles Chemistry Mole Questions And Answers
  - Highlighting and Note-Taking Moles Chemistry Mole Questions And Answers
  - Interactive Elements Moles Chemistry Mole Questions And Answers
8. Staying Engaged with Moles Chemistry Mole Questions And Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Moles Chemistry Mole Questions And Answers
9. Balancing eBooks and Physical Books Moles Chemistry Mole Questions And Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Moles Chemistry Mole Questions And Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Moles Chemistry Mole Questions And Answers
  - Setting Reading Goals Moles Chemistry Mole Questions And Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Moles Chemistry Mole Questions And Answers
  - Fact-Checking eBook Content of Moles Chemistry Mole Questions And Answers
  - Distinguishing Credible Sources
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  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
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  - Integration of Multimedia Elements

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