

**Answer and Chemical Formulas**

**Free-Response Answer Guidelines**

- (a) If answered by describing trends or components:
  - For trends, reference a component that includes both number and identity of ions.
  - Always include the unit charge.
- (b) Reference a charge when describing a trend:
  - In systems to which electronegativity is used, electronegativity is a relative property and trends reflecting electronegativity are relative trends.
- (c) If answering a question in detail:
  - Write a balanced chemical equation, with correct stoichiometric coefficients, for the entire reaction.

**A. Finding the Molar Mass**

- (a)  $\text{C}_2\text{H}_6\text{O}$  (molar mass = 46.07 g)
  - (b)  $\text{C}_2\text{H}_4\text{O}_2$  (molar mass = 60.05 g)
  - (c)  $\text{C}_2\text{H}_4\text{O}_2$  (molar mass = 60.05 g)
  - (d)  $\text{C}_2\text{H}_6\text{O}$  (molar mass = 46.07 g)
  - (e)  $\text{C}_2\text{H}_4\text{O}_2$  (molar mass = 60.05 g)
  - (f)  $\text{C}_2\text{H}_6\text{O}$  (molar mass = 46.07 g)
  - (g)  $\text{C}_2\text{H}_4\text{O}_2$  (molar mass = 60.05 g)
  - (h)  $\text{C}_2\text{H}_6\text{O}$  (molar mass = 46.07 g)
  - (i)  $\text{C}_2\text{H}_4\text{O}_2$  (molar mass = 60.05 g)
- 100 g of  $\text{C}_2\text{H}_6\text{O}$  is 2.17 mol  $\text{C}_2\text{H}_6\text{O}$  (46.07 g  $\text{C}_2\text{H}_6\text{O}$  = 1 mol).**

# Moles And Chemical Formulas Lab Manual

**ML Yell**



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