

Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are *three-dimensional*. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

Model 1 – Lewis Structures

Lewis Structures	H ₂ CO 3 electron domains (3 bonding, 0 nonbonding)	3-D Molecular Shape
<div> <div>1. H₂CO</div> <div> $\begin{array}{c} \text{:O:} \\ \\ \text{H}-\text{C}-\text{H} \end{array}$ </div> </div> <div> <div>2. BeF₂</div> <div> $\begin{array}{c} \text{:F:} \quad \text{:Be:} \quad \text{:F:} \\ \quad \\ \text{:F:} \quad \text{:Be:} \quad \text{:F:} \end{array}$ </div> </div> <div> <div>3. CH₄</div> <div> $\begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \end{array}$ </div> </div>	<div>BeF₂</div> <div>2 electron domains (2 bonding, 0 nonbonding)</div>	
<div> <div>4. NH₃</div> <div> $\begin{array}{c} \text{H} \\ \\ \text{H}-\text{N}-\text{H} \\ \\ \text{:} \end{array}$ </div> </div> <div> <div>5. H₂O</div> <div> $\begin{array}{c} \text{:O:} \\ \\ \text{H}-\text{O}-\text{H} \\ \\ \text{:} \end{array}$ </div> </div> <div> <div>6. CO₂</div> <div> $\begin{array}{c} \text{:O:} \quad \text{:C:} \quad \text{:O:} \\ \quad \\ \text{:O:} \quad \text{:C:} \quad \text{:O:} \end{array}$ </div> </div>	<div>CH₄</div> <div>4 electron domains (4 bonding, 0 nonbonding)</div>	
	<div>NH₃</div> <div>4 electron domains (3 bonding, 1 nonbonding)</div>	
	<div>H₂O</div> <div>4 electron domains (2 bonding, 2 nonbonding)</div>	
	<div>CO₂</div> <div>2 electron domains (2 bonding, 0 nonbonding)</div>	
<div>Lone pair = ••</div>		

Molecular Geometry Pogil Answers Lewis Structures

Arturo Cuomo



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