

Molecular Polarity

How do we know if one molecule will be attracted to another molecule?

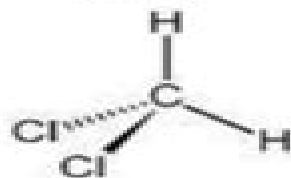
Why?

Did you ever wonder how molecules of a substance stay close together? Why does water form droplets or support things on its surface? Electrons may be distributed evenly or unevenly throughout a molecule, which creates partial charges at different parts of the molecule. These partial charges on one molecule often interact with the partial charges of a neighboring molecule. In this activity you will learn how to determine molecular polarity and placement of partial charges molecules.

Model 1 – The Molecules

POLAR MOLECULES:

Tetrahedral



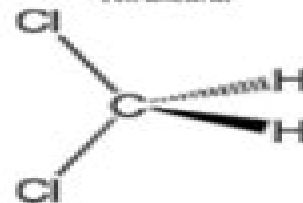
(CH₂Cl₂)

Trigonal Pyramidal



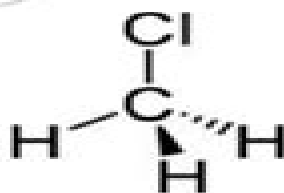
(NH₃)

Tetrahedral



(CH₃Cl)

Tetrahedral



(CH₃Cl)

Bent



(H₂O)

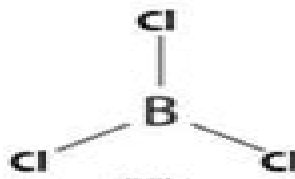
NON-POLAR MOLECULES:

Tetrahedral



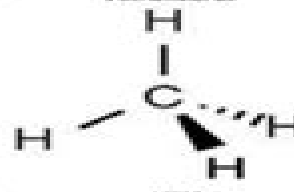
(CCl₄)

Trigonal Planar



(BCl₃)

Tetrahedral



(CH₄)

Linear



(CO₂)

Trigonal Planar



(BH₃)

Bond Type Guide:

- C – Cl = Polar Covalent
- C – H = Polar Covalent
- C – O = Polar Covalent
- N – H = Polar Covalent
- O – H = Polar Covalent
- B – Cl = Polar Covalent
- B – H = Non-Polar Covalent
- B – Al = Non-Polar Covalent

Molecule Polarity Pogil

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