

Name _____

Molecular Geometries Lab Report Sheet

Formula	Lewis Structure	No of electron domains/Atp	Electron domain geometry	VSEPR Molecular geometry	Hybridization of central atom	Polar ?
CO ₂	$\begin{array}{c} \text{Valence } 24 \\ \text{:O}=\text{C}=\text{O:} \\ \text{2 charge clouds} \end{array}$	0	Linear 180°	Linear 	SP	NP
BF ₃	$\begin{array}{c} \text{Valence } 32 \\ \text{:F:} \\ \text{:F}-\text{B}-\text{F:} \end{array}$	0	Trigonal Planar 120°		SP ₂	NP
SnCl ₂	$\begin{array}{c} \text{Valence } 24 \\ \text{:Cl}=\text{Sn}-\text{Cl:} \end{array}$	3	Bent 120°		SP ₂	P
BF ₄ ⁻	$\begin{array}{c} \text{Valence } 32 \\ \left[\begin{array}{c} \text{:F:} \\ \text{:F}-\text{B}-\text{F:} \\ \text{:F:} \end{array} \right]^- \end{array}$	0	Tetrahedral 109.5°		sp ³	P
AsH ₃	$\begin{array}{c} \text{H} \\ \\ \text{H}-\text{As}-\text{H} \end{array}$	4	Trigonal Pyramidal 109.5°		sp ³	P
H ₂ S	$\text{H}-\ddot{\text{S}}-\text{H}$	2	Bent 109.5°		sp ³	NP
NH ₂ ⁻	$\left[\text{H}-\ddot{\text{N}}-\text{H} \right]^-$	2	Bent 109.5°		sp ³	P
POF ₃	$\begin{array}{c} \text{Valence } 40 \\ \text{:F:} \quad \text{:F:} \\ \diagdown \quad \diagup \\ \text{:F}-\text{P}-\text{Cl:} \\ \diagup \quad \diagdown \\ \text{:F:} \end{array}$	0	Trigonal Bipyre 120° + 90°		sp ³ d	P

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Table of Contents Molecular Geometry And Polarity Lab Answers

1. Understanding the eBook Molecular Geometry And Polarity Lab Answers
 - The Rise of Digital Reading Molecular Geometry And Polarity Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Molecular Geometry And Polarity Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molecular Geometry And Polarity Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molecular Geometry And Polarity Lab Answers
 - Personalized Recommendations
 - Molecular Geometry And Polarity Lab Answers User Reviews and Ratings
 - Molecular Geometry And Polarity Lab Answers and Bestseller Lists

5. Accessing Molecular Geometry And Polarity Lab Answers Free and Paid eBooks
 - Molecular Geometry And Polarity Lab Answers Public Domain eBooks
 - Molecular Geometry And Polarity Lab Answers eBook Subscription Services
 - Molecular Geometry And Polarity Lab Answers Budget-Friendly Options
6. Navigating Molecular Geometry And Polarity Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Molecular Geometry And Polarity Lab Answers Compatibility with Devices
 - Molecular Geometry And Polarity Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molecular Geometry And Polarity Lab Answers
 - Highlighting and Note-Taking Molecular Geometry And Polarity Lab Answers
 - Interactive Elements Molecular Geometry And Polarity Lab Answers
8. Staying Engaged with Molecular Geometry And Polarity Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molecular Geometry And Polarity Lab Answers
9. Balancing eBooks and Physical Books Molecular Geometry And Polarity Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molecular Geometry And Polarity Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molecular Geometry And Polarity Lab Answers
 - Setting Reading Goals Molecular Geometry And Polarity Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molecular Geometry And Polarity Lab Answers
 - Fact-Checking eBook Content of Molecular Geometry And Polarity Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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