

Unit 7 Test Review-Answer Key

1. What is the percent composition of Si_2O_6 ?

$$\begin{aligned} \bullet \text{ Si} &= (2 \times 28.09) = 56.18 \div 152.18 \times 100 = 36.9 \% \\ \bullet \text{ O} &= (6 \times 16) = 96 \div 152.18 \times 100 = 63.1 \% \\ \text{Total MM} &= 152.18 \end{aligned}$$

2. What is the percent composition of $\text{Ti}(\text{SO}_4)_2$?

$$\begin{aligned} \text{Ti} &= (1 \times 47.88) = 47.88 \div 240.02 \times 100 = 19.9 \% \\ \text{S} &= (2 \times 32.07) = 64.14 \div 240.02 \times 100 = 26.7 \% \\ \text{O} &= (8 \times 16) = 128 \div 240.02 \times 100 = 53.3 \% \\ \text{Total MM} &= 240.02 \end{aligned}$$

3. What is the percent composition of Hydrogen in $\text{C}_6\text{H}_{12}\text{O}_6$?

$$\begin{aligned} \text{C} &= (6 \times 12.01) = 72.06 \div 180.18 \times 100 \\ \text{H} &= (12 \times 1.01) = 12.12 \div 180.18 \times 100 = 6.7 \% \\ \text{O} &= (6 \times 16) = 96 \div 180.18 \times 100 \\ \text{Total MM} &= 180.18 \end{aligned}$$

4. An unknown compound is 43.4% sodium, 11.3% carbon, and 45.3% oxygen. What is the empirical formula?

$$\text{Na} = 43.4 \text{ g} \times \frac{1 \text{ mole}}{22.99 \text{ g}} = 1.89 \text{ moles} \div 0.94 = 2$$

$$\text{C} = 11.3 \text{ g} \times \frac{1 \text{ mole}}{12.01 \text{ g}} = 0.94 \text{ moles} \div 0.94 = 1 \quad \text{Empirical Formula} = \text{Na}_2\text{CO}_3$$

$$\text{O} = 45.3 \text{ g} \times \frac{1 \text{ mole}}{16 \text{ g}} = 2.93 \text{ moles} \div 0.94 = 3$$

5. Monosodium glutamate (MSG) is sometimes added to food to enhance flavor. Analysis determined this compound to be 35.5% C, 4.77% H, 8.28% N, 13.6% Na, and 37.9% O. What is the empirical formula for MSG?

$$\text{C} = 35.5 \text{ g} \times \frac{1 \text{ mole}}{12.01 \text{ g}} = 2.96 \text{ moles} \div 0.59 = 5.01 = 5$$

$$\text{H} = 4.77 \text{ g} \times \frac{1 \text{ mole}}{1.01 \text{ g}} = 4.7 \text{ moles} \div 0.59 = 7.96 = 8$$

$$\text{N} = 8.28 \text{ g} \times \frac{1 \text{ mole}}{14.01 \text{ g}} = 0.59 \text{ moles} \div 0.59 = 1$$

$$\text{Na} = 13.6 \text{ g} \times \frac{1 \text{ mole}}{22.99 \text{ g}} = 0.59 \text{ moles} \div 0.59 = 1$$

$$\text{O} = 37.9 \text{ g} \times \frac{1 \text{ mole}}{16 \text{ g}} = 2.37 \text{ moles} \div 0.59 = 4.01 = 4$$

$$\text{Empirical Formula} = \text{C}_5\text{H}_8\text{NNaO}_4$$

Mole Test Answer

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