

1 Mole to Mole Relationship between Cu and

Ag_Essay1_Experimental Data Table

Record the following data in the table below.

mass of empty beaker	110.000 g
mass of beaker with silver nitrate (after all additions) and copper	331.634 g
mass of beaker with silver	113.395 g

2 Mole to Mole Relationship between Cu and

Ag_Essay2_Moles Silver Nitrate Added to Reaction

How many moles of silver nitrate were added to the reaction?

$$n = M \times V$$

$$n = 0.2 \text{ M} \times 0.200 \text{ L} = 0.04 \text{ mol AgNO}_3$$

3 Mole to Mole Relationship between Cu and

Ag_Essay3_Silver Ions Consumed

Did all the silver ions get consumed in the reaction? The molar mass of silver is 107.87 g/mol. Justify your answer.

$$n_{\text{Ag}^+} = M \times V$$

$$n_{\text{Ag}^+ \text{ added}} = 0.2 \text{ M} \times 0.200 \text{ L} = 0.04 \text{ mol Ag}^+$$

$$n_{\text{Ag}^+ \text{ consumed}} = \frac{m}{MM} = \frac{113.395 \text{ g} - 110.000 \text{ g}}{107.87 \text{ g/mol}} = 0.03147 \text{ mol Ag}$$

No, not all the silver ions were consumed in the reaction.

4 Mole to Mole Relationship between Cu and

Ag_Essay4_Silver Nitrate Plus Zinc Reaction

Mole Relationship Answers

**Donald J. Wink, Sharon Fetzer-
Gislason, Sheila McNicholas**



Mole Relationship Answers:

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Basic Concepts of Chemistry Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone's *Basic Concepts of Chemistry* provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter WileyPLUS sold separately from text

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