

## THEE NIKELBLAGENACTIVITEIT

### Opdracht 1:

15.4. Bereken de volgende hoeveelheid overblijfsel van eenmaal het afbreken van de hoeveelheid van 1000 atomen van Nippon-1.

|                                                                  |    |      |
|------------------------------------------------------------------|----|------|
| Hoeveel atomen van Nippon-1 overblijft?                          | 10 | 1000 |
| Hoeveel atomen van Nippon-1 overblijft na 10 halveringsperiodes? | 10 | 1000 |
| Hoeveel atomen van Nippon-1 overblijft na 10 halveringsperiodes? | 10 | 1000 |
| Hoeveel atomen van Nippon-1 overblijft na 10 halveringsperiodes? | 10 | 1000 |

16. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

17. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

18. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

19. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

### Opdracht 2:

20. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

21. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

22. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

### Opdracht 3:

23. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

24. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

25. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

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29. 1000 atomen Nippon-1 overblijft na 10 halveringsperiodes. Hoeveel atomen Nippon-1 overblijft na 10 halveringsperiodes?

# Mole Lab Activity Answers

**Albert A Gayle**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located on the right side of the slide.

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Exploring Anatomy in the Laboratory, Second Edition Erin C Amerman, 2021-01-01 This comprehensive beautifully illustrated and affordably priced manual is appropriate for a one semester anatomy only laboratory course The unique interactive approach of these exercises helps students develop a deeper understanding of the material as they prepare to embark on allied health careers Through focused activities and by eliminating redundant exposition and artwork found in most primary textbooks this manual complements the lecture material and serves as an efficient and effective tool for learning in the lab *ChemDiscovery Lab Manual* Olga I. Agapova, 2002

**CLASS 11 CHEMISTRY 5 SOLVED "CASE STUDIES"** Priti Singhal, 2024-11-17 This book is structured to align with the latest syllabus and curriculum guidelines ensuring that the content is both relevant and rigorous Each chapter begins with a clear set of learning objectives providing a roadmap for students to understand what they will achieve by the end of the chapter We have included numerous diagrams illustrations and real life examples to make complex concepts more accessible and engaging *Regents Exams and Answers: Chemistry--Physical Setting Revised Edition* Barron's Educational Series, Albert Tarendash, 2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the Chemistry Regents including actual recently administered exams and thorough answer explanations for all questions This book features Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day *Exploring General Chemistry in the Laboratory* Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

**Stoichiometry Unit Project** Luann Marie Decker, 1998 *Resources for Teaching Middle School Science* Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-04-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them Resources for Teaching Middle School Science developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the

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Analytical and Inorganic Chemistry Laboratory Manual Mr. Rohit Manglik,2024-03-24 Provides practical experiments and procedures in analytical and inorganic chemistry to reinforce theoretical concepts *TID.* ,19?? Reprocessing of Irradiated Fission Reactor Fuel and Breeding Materials ,1964 **The Zinc and Iodine Book** Stephen DeMeo,2019-05-15 This book is for chemistry teachers who are thinking about reinventing their laboratory experiments that they provide to their students More than a collection of experiments it is an example of using a chemical theme to teach chemistry Instead of introducing many different chemicals per experiment as is the norm in most lab manuals this novel resource focuses on two commonly found elements Zinc and Iodine So what is so special about these elements At the heart of this resource is a colorful cyclic reaction between zinc and iodine one that produces a compound that can decompose back to its original elements This unique phenomenon demonstrates that matter not only changes but is also conserved through a chemical reaction Knowing that a compound can be the same but different than the reactants that formed it is to understand the essence of chemical change Complementing this reaction this book contains experimental activities that utilize the zinc and iodine theme to scaffold new concepts such as the properties of matter solid and gas stoichiometry equilibrium kinetics acid

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