

(c) Mass of 1 mole of anhydrous calcium chloride (CaCl_2) = $40 + 71 = 111 \text{ g}$

But 1 molecule of CaCl_2 contains 2 chloride ions₂₃
 $\therefore 111 \text{ g}$ of anhydrous CaCl_2 contains $2 \times 6.023 \times 10^{23}$
 $= 12 \times 10^{23}$ chloride ions

Q2) How many moles of NaOH are present in 160 g of it?

Soln: Mass of 1 mole of $\text{NaOH} = 23 + 16 + 1 = 40 \text{ g}$

40 g of NaOH contains 6×10^{23} no. of molecules

Moles in 160 g of $\text{NaOH} = \frac{160}{40} = \underline{4 \text{ moles}}$

Q3) What is the mass in grams of 2.5 moles of slaked lime Ca(OH)_2 .

Soln: Mass of 1 mole of slaked lime $\text{Ca(OH)}_2 = 40 + 32 + 2$
 $= 74 \text{ g}$

So mass of 2.5 moles of $\text{Ca(OH)}_2 = 2.5 \times 74 = \underline{185 \text{ g}}$

Means 2.5 moles of Ca(OH)_2 weigh 185 g.

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