

Stoichiometry and The Mole – Review Guide

Remember:

- 1 mole = 6.02×10^{23} atoms, molecules, formula units
- 1 mole = grams (of molar mass)
- 1 mole = 22.4 liters (L) of gas



1) The process used to predict the amount of a product that will be formed is stoichiometry.

2) If you know the mass of a substance, you can find the number of moles by dividing it by the molar mass of substance.

3) If you know the number of particles (atoms) of a substance, you can find the number of moles by dividing it by Avogadro's Number.

4) Summarize the steps of a gram-to-gram stoichiometry problem:

Grams A \rightarrow Moles A \rightarrow Moles B \rightarrow Grams B

5) The essential stoichiometry problem needed to convert from one chemical to another is:
Mole to Mole

6) Two reactants undergo a chemical reaction. At the end of the reaction, Chemical A still exists in the container, while Chemical B has been completely used up. Which is the limiting reactant?
Chemical B

7) What is a mole? A unit of measurement used to count atoms/molecules.

8) What is Avogadro's Number? 6.02×10^{23}

9) What does Avogadro's number tell us? The number of atoms or molecules in a mole.

Mole Study Guide Answer Key

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Katherine**



Mole Study Guide Answer Key:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Organic Chemistry Study Guide* Robert J. Ouellette, J. David Rawn, 2014-11-04 Organic Chemistry Study Guide Key Concepts Problems and Solutions features hundreds of problems from the companion book Organic Chemistry and includes solutions for every problem Key concept summaries reinforce critical material from the primary book and enhance mastery of this complex subject Organic chemistry is a constantly evolving field that has great relevance for all scientists not just chemists For chemical engineers understanding the properties of organic molecules and how reactions occur is critically important to understanding the processes in an industrial plant For biologists and health professionals it is essential because nearly all of biochemistry springs from organic chemistry Additionally all scientists can benefit from improved critical thinking and problem solving skills that are developed from the study of organic chemistry Organic chemistry like any skill is best learned by doing It is difficult to learn by rote memorization and true understanding comes only from concentrated reading and working as many problems as possible In fact problem sets are the best way to ensure that concepts are not only well understood but can also be applied to real world problems in the work place Helps readers learn to categorize analyze and solve organic chemistry problems at all levels of difficulty Hundreds of fully worked practice problems all with solutions Key concept summaries for every chapter reinforces core content from the companion book

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