Molar Ratio Practice Problems

Following each equation are two requests for molar ratios from the equation.

1) $N_2 + 3 H_2 \rightarrow 2 NH_3$

N₂ to H₂ NH₃ to H₂

2) 2 SO₂ + O₂ - 2 SO₃

O2 to SO3 O2 to SO2

3) PCl₃ + Cl₂ → PCl₅

PCl₃ to Cl₂ PCl₃ to PCl₅

4) 4 NH₃ + 3 O₂ → 2 N₂ + 6 H₂O

NH₃ to N₂ H₂O to O₂

5) Fe₂O₃ + 3 CO → 2 Fe + 3 CO₂

CO to CO₃ Fe to CO

Mole to Mole Practice Problems

Here's the equation to use for all three problems.

1) How many moles of H₂O are produced when 5 moles of oxygen are used?

Remember: Starting with five moles of oxygen and based on the balanced equation, for every: I mole of oxygen used, two moles of water are produced.

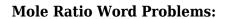
2) If 3.00 moles of H₂O are produced, how many moles of oxygen must be consumed?

3) If 2.5 moles of H2O are produced, how many moles of hydrogen gas must be used,?

4) Suppose 4.00 grams of H₂ were used? How many grams of water would be produced?

Mole Ratio Word Problems

Léo Pomar



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