## MOLARITY BY DILUTION

79+52 = 162 Name \_\_\_\_\_

Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid =  $M \times V$  then,  $M_1 \times V_2 = M_2 \times V_3$ . Solve the following problems.

 How much concentrated 18 M sulfuric acid is needed to prepare 250 mL of a 6.0 M solution?

83 mL]

 How much concentrated 12 M hydrochloric acid is needed to prepare 100,mL of a 2.0 M solution?

- 17mL/
- To what volume should 25 mL of 15 M nitric acid be diluted to prepare a 3.0 M solution?

4. To how much water should 50, mL of 12 M hydrochloric acid be added to produce a 4.0 M solution?

100ml H20)

5. To how much water should 100. mL of 18 M sulfuric acid be added to prepare a 1.5 M solution?

# **Molarity Worksheet With Answers Instructional Fair**

**Guillaume Favre** 

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