

# Dilutions – Practice Problems

**Instructions:** Use the equation for dilutions to complete the problems below.

- 1) A student obtains 0.30 liters of 2.0 M silver nitrate solution. She dilutes the solution until it has a volume of 2.6 liters. What is the **molarity** of the new solution?

$$M_1 V_1 = M_2 V_2$$

$$(2.0) (0.30) = M_2 (2.6)$$

$$(0.6) = M_2 (2.6)$$

$$M_2 = 0.23 \text{ M}$$

- 2) Before an experiment, a scientist obtains 2.45 L of 1.25 M sulfuric acid ( $\text{H}_2\text{SO}_4$ ) and dilutes it until the molarity drops to 0.80 M. What is the **volume** of the new solution?

$$M_1 V_1 = M_2 V_2$$

$$(1.25) (2.45) = (0.8) V_2$$

$$(3.06) = (0.8) V_2$$

$$V_2 = 3.83 \text{ L}$$

- 3) A solution of hydrochloric acid is diluted, bringing the volume from 2.0 L to 8.5 L. If the solution has a final molarity of 1.4 M, what was the **molarity before** dilution?

$$M_1 V_1 = M_2 V_2$$

$$M_1 (2.0) = (1.4) (8.5)$$

$$(250.0) = (2.0) V_2$$

$$V_2 = 125.0 \text{ mL}$$

- 4) If a student adds 250 mL of water to a 300 mL solution of NaCl that has a concentration of 3.0 M, what will be the **molarity** of the final solution?

$$M_1 V_1 = M_2 V_2$$

$$(3.0) (300) = M_2 (550)$$

$$(900) = M_2 (550)$$

$$M_2 = 1.64 \text{ M}$$



- 5) A scientist adds 1.2 liters of water to a 5.0 liter solution, causing the molarity to become 0.4 M. What was the **molarity** of the **initial solution** before the dilution?

$$M_1 V_1 = M_2 V_2$$

$$M_1 (5.0) = (0.4) (7.2)$$

$$M_1 (5.0) = 2.88$$

$$M_1 = 0.576 \text{ M}$$

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