Dilutions - Practice Problems

Instructions: Use the equation for dilutions to complete the problems below.

 A student obtains 0.30 liters of 2.0 M silver mitrate solution. She dilutes the solution until it has a volume of 2.6 liters. What is the molarity of the new solution?

$$M_1V_1 = M_2V_2$$

(2.0) (0.30) = M_2 (2.6)

$$(0.6) = M_2(2.6)$$

 $M_2 = 0.23 M$

2) Before an experiment, a scientist obtains 2.45 L of 1.25 M sulfuric acid (H₂SO₄) and dilutes it until the molarity drops to 0.80 M. What is the volume of the new solution?

$$M_1V_1 = M_2V_2$$

(1.25) (2.45) = (0.8) V_2

$$(3.06) = (0.8) V_2$$

 $V_2 = 3.83 L$

3) A solution of hydrochloric acid is diluted, bringing the volume from 2.0 L to 8.5 L. If the solution has a final molarity of 1.4 M, what was the reclarity before dilution?

$$M_1V_1 = M_2V_2$$

 $M_1(2.0) = (1.4)(8.5)$

$$(250.0) = (2.0) V_2$$

 $V_2 = 125.0 \text{ mL}$

4) If a student adds 250 mL of water to a 300 mL solution of NaC that has a concentration of 3.0 M, what will be the molarity of the final solution?

$$M_1V_1 = M_2V_2$$

(3.0) (300) = M_2 (550)

$$(900) = M_2 (550)$$

 $M_2 = 1.64 M$

5) A scientist adds 1.2 liters of water to a 5.0 liter solution, causing the molarity to become 6.4 M. What was the molarity of the initial solution before the dilution?

$$M_1V_1 = M_2V_2$$

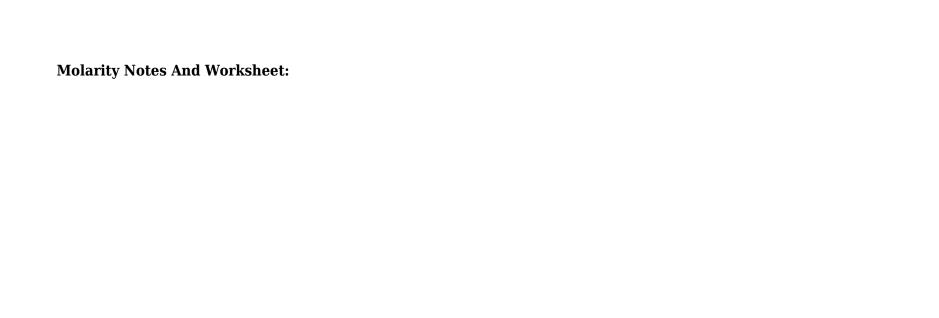
 $M_1(5.0) = (0.4)(7.2)$

$$M_1(5.0) = 2.88$$

$$M_1 = 0.576 M$$

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