Dilutions and Molarity

Dilution is the process used to make a new solution that is less concentrated than the original solution by adding more solvent.

$$\mathbf{M_1V_1} = \mathbf{M_2V_2}$$

M₂ = Final Molarity

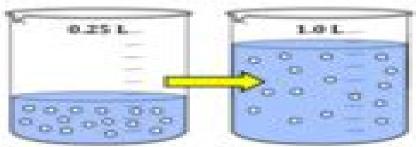
V₂ = Final Volume

Examples Ouestions:

 A scientist obtains 0.25 liters of concentrated 5.0 M NaCl solution. The scientist then dilutes the solution by adding water until the solution has a volume of 1.0 liters. What is the molarity of the solution?

$$M_1 V_1 = M_2 V_2$$

 $(5.0) (0.25) = M_2 (1.0)$
 $(1.25) = M_2 (1.0)$
 $M_2 = 1.25 \text{ M or mol/L}$



2) A concentrated solution of 10.0 M hydrochloric acid solution must be diluted to make it weaker. If a scientist starts with 1.5 liters of the original solution and adds 2.5 liters of water, what will be the final molarity of the solution?

$$M_1V_1 = M_2V_2$$

(10.0) (1.5) = M_2 (4.0)

$$(15.0) = M_2(4.0)$$

 $M_2 = 3.75 \text{ M or mol/L}$

3) In a laboratory, a scientist takes 50 mi, of 5.0 M NaOH solution and adds water until the solution has a melarity of 2.0 M. What is the volume of the new solution?

$$M_1V_1 = M_2V_2$$

(5.0) (50) = (2.0) V_2

$$(250.0) = (2.0) V_2$$

 $V_2 = 125.0 \text{ mL}$

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Richard A. McPherson, Matthew R. Pincus

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