

Name _____ Date _____ Class _____

CHAPTER 16 REVIEW*Reaction Energy***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. For elements in their standard state, the value of ΔH_f° is _____.
2. The formation and decomposition of water can be represented by the following thermochemical equations:



- _____ a. Is energy being taken in or is it being released as liquid H_2O decomposes?
- _____ b. What is the appropriate sign for the enthalpy change in this decomposition reaction?

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

3. _____ If 200. g of water at 20°C absorbs 41 840 J of energy, what will its final temperature be?
4. _____ Aluminum has a specific heat of $0.900 \text{ J/(g}\cdot^\circ\text{C)}$. How much energy in kJ is needed to raise the temperature of a 625 g block of aluminum from 30.7°C to 82.1°C ?
5. The products in a reaction have an enthalpy of 458 kJ/mol, and the reactants have an enthalpy of 658 kJ/mol.
_____ a. What is the value of ΔH for this reaction?

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