



- A) No change occurs
 B) In order to restore equilibrium, the reaction shifts left, toward reactants
 C) In order to restore equilibrium, the reaction shifts right, toward products
 D) In order to restore equilibrium, the reaction shifts right, toward reactants

11. When the temperature is decreased on the following system at equilibrium:
 $2 \text{HCl}(\text{aq}) + \text{Mg}(\text{s}) \rightleftharpoons \text{MgCl}_2(\text{aq}) + \text{H}_2(\text{g}) + \text{heat}$

- A) No change occurs
 B) In order to restore equilibrium, the reaction shifts left, toward reactants
 C) In order to restore equilibrium, the reaction shifts right, toward products
 D) In order to restore equilibrium, the reaction shifts right, toward reactants

12. Consider the following chemical reaction: $\text{H}_{2(g)} + \text{I}_{2(g)} \rightarrow 2 \text{HI}_{(g)}$

At equilibrium in a particular experiment, the concentrations of H_2 , I_2 , and HI were 0.35 M, 0.033 M and 0.55 M respectively. The value of K_{eq} for this reaction is _____.

- A) 23 B) 111 C) 9.0×10^3 D) 6.1 E) 62

13. A reaction vessel is charged with hydrogen iodide, which partially decomposes to molecular hydrogen and iodine: $2 \text{HI}_{(g)} \rightarrow \text{H}_{2(g)} + \text{I}_{2(g)}$. When the system comes to equilibrium at $+25^\circ\text{C}$, $P_{\text{I}_2} = 0.703 \text{ atm}$, and $P_{\text{H}_2} = P_{\text{I}_2} = 0.090 \text{ atm}$. The value of K_p at this temperature is:

- A) 680×10^3 B) 1.30×10^3 C) 1.84×10^3 D) 54.3

E) K_p cannot be calculated for this gas reaction when the volume of the reaction vessel is not given.

14. The value of K_{eq} for the following reaction is 0.25:



The value of K_{eq} at the same temperature for the reaction below is _____.



- A) 0.50 B) 0.063 C) 0.12 D) 0.25 E) 1.6

15. Nitrosyl bromide decomposes according to the following equation:



A sample of NOBr (0.64 mol) was placed in a 1.00-L flask containing no NO or Br_2 . At equilibrium the flask contained 0.46 of NOBr . How many moles of NO and Br_2 , respectively, are in the flask at equilibrium?

- A) 0.18, 0.18 B) 0.46, 0.23 C) 0.18, 0.090
 D) 0.18, 0.360 E) 0.46, 0.46

Short Answer:

1. Complete the following table, based on the following equilibrium system: (6T)
 $\text{SO}_{2(g)} + \text{NO}_{2(g)} \rightleftharpoons \text{SO}_{3(g)} + \text{NO}_{(g)}$ $\Delta H = +460 \text{ kJ}$

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