

Atomic Theory Worksheet

Name: _____

John Dalton (1766 – 1844):

John Dalton was an English chemist. His ideas form the atomic theory of matter. Here are his ideas.

Atomic Theory



- All elements are composed (made up) of atoms. It is impossible to divide or destroy an atom.
- All atoms of the same elements are alike. (One atom of oxygen is like another atom of oxygen.)
- Atoms of different elements are different. (An atom of oxygen is different from an atom of hydrogen.)
- Atoms of different elements combine to form a compound. These atoms must be in definite whole number ratios. For example, water is a compound made up of 2 atoms of hydrogen and 1 atom of oxygen (a ratio of 2:1). Three atoms of hydrogen and 2 atoms of oxygen cannot combine to make water.



1803

1. What is the name of John Dalton's theory? _____
2. What are elements made of? _____
3. An atom of hydrogen and an atom of carbon are _____.
4. What are compounds made of? _____
5. The ratio of atoms in HCl is: a) 1:3 b) 2:1 c) 1:1

J. J. Thompson (Late 1800s):

J. J. Thompson was an English scientist. He discovered the electron when he was experimenting with gas discharge tubes. He noticed a movement in a tube. He called the movement cathode rays. The rays moved from the negative end of the tube to the positive end. He realized that the rays were made of negatively charged particles – electrons.

Atomic Structure Model



J.J. Thompson



1904

1. What did J.J. Thompson discover? _____
2. What is the charge of an electron? _____
3. What are cathode rays made of? _____
4. Why do electrons move from the negative end of the tube to the positive end?

5. What was Thompson working with when he discovered the cathode rays?

NUCLEUS MODEL



ERNEST RUTHERFORD



1911

Lord Ernest Rutherford (1871 – 1937):

Ernest Rutherford conducted a famous experiment called the gold foil experiment. He used a thin sheet of gold foil. He also used special equipment to shoot alpha particles (positively charged particles) at the gold foil. Most particles passed straight through the foil like the foil was not there. Some particles went straight back or were deflected (went in another direction) as if they had hit something. The experiment shows:

- Atoms are made of a small positive nucleus; positive nucleus repels (pushes away) positive alpha particles
- Atoms are mostly empty space

1. What is the charge of an alpha particle? _____
2. How did he know that an atom was mostly empty space?

3. What happened to the alpha particles as they hit the gold foil?

4. How did he know that the nucleus was positively charged?

Modern Atomic Theory Concept Review Answers

Huangqi Zhang



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