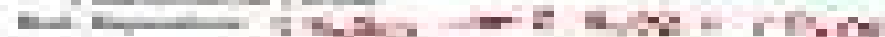


Stoichiometry Worksheet 2: Percent Yield

What aspects of the procedure do you think:

- affected the theoretical or actual yield negatively?
- substantially affected your yield in a positive or negative way? explain how these things can be made up.
- affected your yield in a positive or negative way? explain why these things can be made up.

1. Using the following procedure, calculate the theoretical yield of product. Then, calculate the percent yield of the product. What is the percent yield of the product? (Show your work.)



$$\text{Theoretical yield} = \frac{\text{mass of product}}{\text{mass of reactant}} \times 100\%$$

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$$

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$$

2. Using the following procedure, calculate the theoretical yield of product. Then, calculate the percent yield of the product. What is the percent yield of the product? (Show your work.)



$$\text{Theoretical yield} = \frac{\text{mass of product}}{\text{mass of reactant}} \times 100\%$$

3. Using the following procedure, calculate the theoretical yield of product. Then, calculate the percent yield of the product. What is the percent yield of the product? (Show your work.)



$$\text{Theoretical yield} = \frac{\text{mass of product}}{\text{mass of reactant}} \times 100\%$$

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$$

4. Using the following procedure, calculate the theoretical yield of product. Then, calculate the percent yield of the product. What is the percent yield of the product? (Show your work.)



$$\text{Theoretical yield} = \frac{\text{mass of product}}{\text{mass of reactant}} \times 100\%$$

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$$

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Xiang Xie



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