

## AP Chemistry Review

### Chapter 6 and 7, Atomic Structure and Periodicity

- Be familiar with the electromagnetic spectrum
  1. What types of light are shorter in length than microwave? Know the whole order of the spectrum.
  2. What type of light helps scientists to learn about electron structure?
  3. What creates the emission lines in a spectra for an element? What would cause only certain wavelengths of light to be absorbed by an atom?
- Convert between frequency and wavelength
  4. Calculate the frequency of light emitted by exciting Sr atoms if the wavelength is 461 nm.
  5. Calculate the wavelength of light emitted by exciting Na atoms if the frequency is  $3.45 \times 10^{14} \text{ s}^{-1}$
- Calculate the energy of a wave or an electron
  6. Calculate the energy of 1 mole of photons emitted by a Li atom if the wavelength is 670nm.
- Define and understand:
  7. Ground state, excited state, Pauli exclusion principle, Heisenberg Uncertainty Principle, Hund's rule.
- Know the maximum number of electrons in and shape of each type of orbital.
  8. 3s                      5p                      4f                      6d
- Following the Aufbau principle, determine electron configurations, both longhand and condensed, for elements and ions
- Match each explanation below with the correct letter
  - (A)  $1s^2 2s^2 2p^5 3s^2 3p^5$
  - (B)  $1s^2 2s^2 2p^6 3s^2 3p^6$
  - (C)  $1s^2 2s^2 2p^6 2d^{10} 3s^2 3p^6$
  - (D)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5$
  - (E)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^2$
- 9. An impossible electron configuration
- 10. The ground-state configuration for the atoms of a transition element
- 11. The ground-state configuration of a negative ion of a halogen
- 12. The ground-state configuration of a common ion of an alkaline earth element
- 13. List 3 species that are isoelectronic
- 14. Which of the following represents the ground state electron configuration for the  $\text{Mn}^{3+}$  ion? (Atomic number Mn = 25)
  - (a)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^4$
  - (b)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^2$
  - (c)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2 4s^2$
  - (d)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$
  - (e)  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^1$
- Draw Dot diagrams for elements
  15. Boron                      Calcium                      Krypton

# Merrill Chemistry Study Guide Chapter 6 Answers

**TD Snyder**



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