

Everyday Mathematics

Grade 1

OA = Oral Assessment
SA = Slate Assessment
AA = Alternative Assessment
WA = Written Assessment

Unit 1

Establishing Routines

Time to complete: 18 days

Secure Goals	GLEs	Assessment
Count up and back by 1s, starting with any number up to and including 20		OA 1,2; WA 3
Count up to 20 objects		SA 6,7; WA 3
Developing/Secure Goals	GLEs	Assessment
Developing Goals	GLEs	Assessment
Count by 5s to 40		OA 3
Count by 2s to 40		OA 4
Write numbers from 1 to 20		SA 1-7; WA 1, 3, 6
Compare pairs of numbers less than 16		SA 1-5; WA 3-5
Write and count tallies		WA 1,2
Beginning/Developing Goals	GLEs	Assessment
Beginning Goals	GLEs	Assessment

Unit 2

Everyday Uses of Numbers

Time to complete: 18 days

Secure Goals	GLEs	Assessment
Count by 2s to 20; count by 5s to 50		OA 1-3
Developing/Secure Goals	GLEs	Assessment
Count up and back by 1s on the number grid		OA 5; SA 1
Exchange pennies for nickels		WA 5
Tell time to the nearest hour		OA 6; WA 1,2
Developing Goals	GLEs	Assessment

Maths Scope For 2014 Final For Grade1

Terry C. Jones



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$$2\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$
 $8.3 \text{ mole } \text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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 $\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles } \text{O}_2$

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 $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 b determine the theoretical yield of Al_2O_3
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g/mol}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 150.1 \text{ g } \text{Al}_2\text{O}_3$
 $302.3 \text{ g } \text{Al}_2\text{O}_3$
 $26.98 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g/mol}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 25.4 \text{ g } \text{Al}_2\text{O}_3$
 c determine the percent yield
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$
 theoretical 302.3 g

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