

Thermodynamics

Section Quiz: The First Law of Thermodynamics

Write the letter of the correct answer in the space provided.

- _____ 1. Which concept does the first law of thermodynamics describe?
 - a. conservation of mass
 - b. conservation of energy
 - c. energy being transformed
 - d. conservation of momentum
- _____ 2. Which concept allows $Q = 0$ for the first law of thermodynamics when the heat of $Q_1 = -Q_2$?
 - a. an isolated system process
 - b. an adiabatic process
 - c. an isobaric process
 - d. an isochoric process
- _____ 3. Which concept allows $W = 0$ for the first law of thermodynamics when the heat of $Q_1 = Q_2$?
 - a. an isobaric process
 - b. an isochoric process
 - c. an isobaric process
 - d. an isochoric process
- _____ 4. Which process is isothermal, $Q = W = 0$, so that the first law of thermodynamics takes the form of $Q_1 = Q_2$?
 - a. an isobaric process
 - b. an isochoric process
 - c. an isobaric process
 - d. an isochoric process
- _____ 5. Which concept allows $Q_1 = 0$ so that the first law of thermodynamics takes the form of $Q_2 = W$?
 - a. an isobaric process
 - b. an isochoric process
 - c. an isobaric process
 - d. an isochoric process
- _____ 6. A process might follow the $Q_1 = 0$ of energy is heat, giving up $Q_2 = 0$ of energy as heat, the surroundings, and some $W = 0$ of work. What further change is the internal energy of the system?
 - a. zero J
 - b. 1000 J
 - c. 2000 J
 - d. 3000 J

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