

CHAPTER 11 REVIEW

Ions in Aqueous Solutions and Colligative Properties

SECTION REVIEW

IMPORT SKILLS Answer the following questions in the space provided.

1. Identify the ions represented under each of the descriptions on the left.

- | | |
|--|-------------------------------------|
| <input type="text"/> an ionic compound that is quite soluble in water | (a) NaCl |
| <input type="text"/> an ionic compound that is not very soluble in water | (b) Na_2SO_4 |
| <input type="text"/> a molecular compound that contains no oxygen | (c) AgCl |
| <input type="text"/> a molecular compound that does not contain hydrogen | (d) $\text{C}_2\text{H}_5\text{OH}$ |

2. Consider separately separately-saturated solutions of various liquid solutes in water. In each case, the following statements:

- | | |
|----------------------|--|
| <input type="text"/> | a. The change in the boiling point decreases with the density of the solid solute, with a constant volume fraction of solute. |
| <input type="text"/> | b. The change in the boiling point decreases with the density of the solid solute, with a constant volume fraction of solute. |
| <input type="text"/> | c. The change in the boiling point becomes greater as the concentration of the solute in solution increases decreases. |

3. a. Explain why compound is soluble in water. Consider whether the formation of particles is favorable or unfavorable.

b. Identify any species that in the system give distinctive gases.

c. Write the net ionic equation for the formation of water vapor.

4. Explain why applying road salt helps melt ice and why it is not a good idea to use it in large quantities.

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