

Electrochemistry: Unit Review

- What is the oxidation number of chlorine in HClO_4 ?
 1. -1 2. +1 3. +7 4. +8
- What is the oxidation number of carbon in C_2H_6 ?
 1. -6 2. -3 3. +3 4. +6
- A battery is an example of:
 1. An electrochemical cell 2. A electrolytic cell 3. A pressure cell
- Given the reaction: $\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$ Which equation represents the oxidation that takes place?
 1. $\text{Mg}^{+2} + 2e^- \rightarrow \text{Mg}$ 2. $\text{Mg} \rightarrow \text{Mg}^{+2} + 2e^-$
 3. $\text{Cu}^{+2} + 2e^- \rightarrow \text{Cu}$ 4. $\text{Cu} \rightarrow \text{Cu}^{+2} + 2e^-$
- In the following reaction: $\text{Fe} + \text{Zn}^{+2} \rightarrow \text{Fe}^{+2} + \text{Zn}$
 A. Fe is reduced and Zn^{+2} is the oxidizing agent
 B. Fe is reduced and Zn^{+2} is the reducing agent
 C. Fe is oxidized and Zn^{+2} is the oxidizing agent
 D. Fe is oxidized and Zn^{+2} is the reducing agent
- In which species does oxygen have an oxidation number of -1?
 A. H_2O 2. Na_2O_2 3. Al_2O_3 4. O_2
- Elements that readily gain electrons tend to have
 A. High ionization energies and high electronegativities
 B. High ionization energies and low electronegativities
 C. Low ionization energies and high electronegativities
 D. Low ionization energies and low electronegativities
- In an electrochemical cell, the cathode is
 1. neutral 2. negative 3. positive
- What is the purpose of a salt bridge in an electrochemical cell?
 1. To act as a reducing agent 2. To act as an oxidizing agent
 3. To provide electrons that flow between the two half cells and complete the electrical circuit
 4. To provide ions that flow between the two half cells and complete the electrical circuit
- What is NOT the same for an electrochemical cell and an electrolytic cell?
 A. An oxidation/reduction chemical reaction occurs within them
 B. They have two electrodes: a cathode and an anode
 C. They produce voltage and an electric current
 D. Electrons flow through an external circuit
- Given the equations A, B, C, and D; which two equations represent redox reactions?
 A. $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$
 B. $\text{Cl}_2 + \text{H}_2\text{O} \rightarrow \text{HClO} + \text{HCl}$
 C. $\text{CuO} + \text{CO} \rightarrow \text{CO}_2 + \text{Cu}$
 D. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

Holt Electrochemistry Review Answers

Daniel F McAuley



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